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Surname					Other names				
Centre Number					Candidate Number				
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Edexcel GCSE

Chemistry/Science

Unit C1: Chemistry in Our World

Higher Tier

Monday 21 May 2012 – Morning Time: 1 hour	Paper Reference 5CH1H/01
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You must have: Calculator, ruler	Total Marks
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Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*

Information

- The total mark for this paper is 60.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*
- Questions labelled with an **asterisk** (*) are ones where the quality of your written communication will be assessed
– *you should take particular care with your spelling, punctuation and grammar, as well as the clarity of expression, on these questions.*

Advice

- Read each question carefully before you start to answer it.
- Keep an eye on the time.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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PEARSON

The Periodic Table of the Elements

1	2	3	4	5	6	7	0										
7 Li lithium 3	9 Be beryllium 4	11 Na sodium 11	12 C carbon 6	13 Al aluminium 13	14 N nitrogen 7	15 P phosphorus 15	16 O oxygen 8	17 F fluorine 9	18 Ne neon 10								
19 K potassium 19	20 Ca calcium 20	23 Sc scandium 21	24 Ti titanium 22	25 V vanadium 23	26 Cr chromium 24	27 Mn manganese 25	28 Fe iron 26	29 Co cobalt 27	30 Ni nickel 28	31 Cu copper 29	32 Zn zinc 30	33 Ga gallium 31	34 Ge germanium 32	35 As arsenic 33	36 Se selenium 34	37 Br bromine 35	38 Kr krypton 36
37 Rb rubidium 37	38 Sr strontium 38	39 Y yttrium 39	40 Zr zirconium 40	41 Nb niobium 41	42 Mo molybdenum 42	43 Tc technetium 43	44 Ru ruthenium 44	45 Rh rhodium 45	46 Pd palladium 46	47 Ag silver 47	48 Cd cadmium 48	49 In indium 49	50 Sn tin 50	51 Sb antimony 51	52 Te tellurium 52	53 I iodine 53	54 Xe xenon 54
55 Cs caesium 55	56 Ba barium 56	57 La* lanthanum 57	72 Hf hafnium 72	73 Ta tantalum 73	74 W tungsten 74	75 Re rhenium 75	76 Os osmium 76	77 Ir iridium 77	78 Pt platinum 78	79 Au gold 79	80 Hg mercury 80	81 Tl thallium 81	82 Pb lead 82	83 Bi bismuth 83	84 Po polonium 84	85 At astatine 85	86 Rn radon 86
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[277] Hs hassium 108	[268] Mt meitnerium 109	[271] Ds darmstadtium 110	[272] Rg roentgenium 111	Elements with atomic numbers 112-116 have been reported but not fully authenticated						

1 H hydrogen 1

relative atomic mass
atomic symbol
name
atomic (proton) number

* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.



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Questions begin on next page.



Answer ALL questions

Some questions must be answered with a cross . If you change your mind about an answer, put a line through the box and then mark your new answer with a cross .

The reactivity series

1 The list shows some metals in reactivity series order with the most reactive at the top.

most reactive	calcium
	magnesium
	aluminium
	zinc
	iron
	copper
least reactive	gold

(a) Which of these metals can be found as the uncombined metal in the Earth's crust?

Put a cross () in the box next to your answer.

(1)

- A** calcium
- B** gold
- C** magnesium
- D** zinc

(b) Metals are extracted by the reduction of their ores.

State the meaning of the term **reduction**.

(1)

.....

.....



(c) Aluminium is extracted by the electrolysis of a molten mixture of its ore (bauxite) and cryolite.

Iron is extracted by heating a mixture of its ore and carbon.

Explain why electrolysis is used to extract aluminium but is not used to extract iron.

(2)

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(d) An alloy of aluminium with magnesium is used for parts of aeroplanes.



Explain why the aluminium alloy is stronger than pure aluminium.

(2)

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(e) When aluminium corrodes, it reacts with oxygen to form aluminium oxide, Al_2O_3 .

Complete the balancing of this equation by putting numbers in the spaces provided.

(2)



(Total for Question 1 = 8 marks)



Carbon dioxide levels in the atmosphere

2 Scientists believe that, about 4500 million years ago, the Earth was very much hotter than it is now.

The atmosphere then was very different from that on Earth today.

(a) Explain why it is difficult to be precise about the composition of the Earth's early atmosphere.

(2)

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(b) As the Earth cooled, oceans formed.

How did this affect the composition of the atmosphere?

(1)

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(c) The first plants appeared about 400 million years ago.

Explain how the growth of these plants affected the composition of the atmosphere.

(2)

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(d) What is the current approximate percentage of carbon dioxide in the Earth's atmosphere?

Put a cross (☒) in the box next to your answer.

(1)

- A** 0.04
- B** 1.0
- C** 4.0
- D** 10



- (e) The amount of carbon dioxide in the Earth's atmosphere has been rising over the past fifty years, mainly caused by an increase in the quantity of fossil fuels that have been burned.

State another cause of increasing amounts of carbon dioxide in the atmosphere.

(1)

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.....
(Total for Question 2 = 7 marks)



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Hydrocarbons

- 3 (a) During fractional distillation, crude oil is separated into a number of fractions.

The table shows the relative amount of these fractions that can be obtained from crude oil.

The table also shows the relative demand for each of these fractions.

fraction	relative amount obtained	relative demand
LPG	2	6
petrol	12	29
kerosene	16	11
diesel	24	29
fuel oil and bitumen	46	25

- (i) For which fractions does the demand exceed the supply?

Put a cross (☒) in the box next to your answer.

(1)

- A** kerosene, diesel, fuel oil and bitumen
- B** LPG, petrol and diesel
- C** LPG, petrol and kerosene
- D** petrol, diesel, fuel oil and bitumen

- (ii) In another process, called cracking, large molecules in some fractions are converted into smaller molecules.

Explain why cracking is needed.

(2)

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(b) The table shows the number of carbon atoms per molecule in the substances present in each of the fractions.

fraction	number of carbon atoms per molecule
LPG	1 – 4
petrol	4 – 12
kerosene	9 – 16
diesel	15 – 25
fuel oil and bitumen	over 25

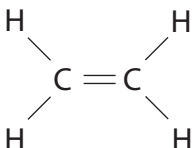
Complete the sentence by putting a cross (☒) in the box next to your answer.

(1)

As the number of carbon atoms in the molecules of a hydrocarbon increases

- A** the number of hydrogen atoms in the molecule remains the same
- B** the boiling point of the hydrocarbon increases
- C** the hydrocarbon becomes easier to burn
- D** the viscosity of the hydrocarbon decreases

(c) The structure of a molecule of ethene is



(i) What is the formula of a molecule of ethene?

Put a cross (☒) in the box next to your answer.

(1)

- A** CH₂
- B** C₂H₄
- C** C_nH_{2n}
- D** (CH₂)_n



(ii) Ethene can be converted into poly(ethene).

Write a balanced equation for this reaction.

(2)

(d) Many power stations generate electricity by burning fossil fuels, such as fuel oil. This process adds carbon dioxide to the atmosphere.

(i) Explain why some people are concerned about the increase in the amount of carbon dioxide in the atmosphere.

(2)

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(ii) Some hydrocarbon fuels can contain sulfur impurities.

Explain how the product of combustion of these sulfur impurities affects the environment.

(2)

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(Total for Question 3 = 11 marks)



Calcium compounds

- 4 Calcium carbonate is an important raw material in the chemical industry. It exists naturally as chalk, marble and limestone.

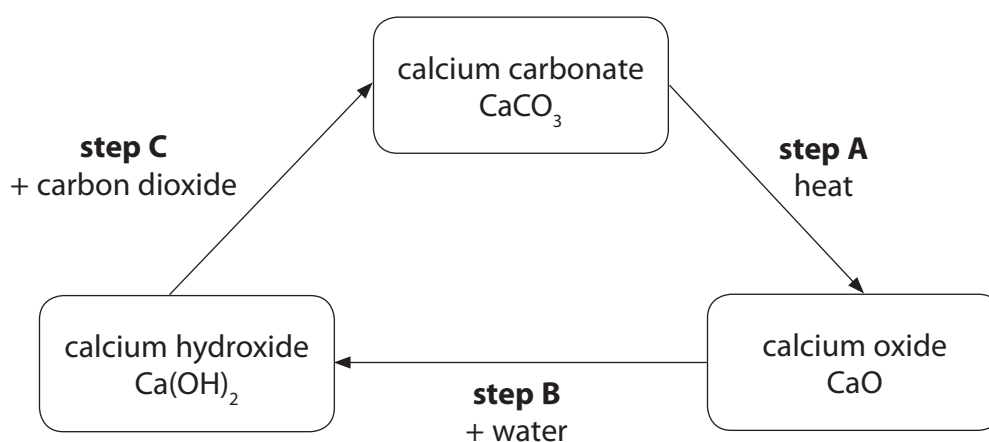
(a) Which row of the table correctly identifies chalk, marble and limestone as sedimentary or metamorphic rocks?

Put a cross (☒) in the box next to your answer.

(1)

		chalk	marble	limestone
<input type="checkbox"/>	A	sedimentary	sedimentary	metamorphic
<input type="checkbox"/>	B	metamorphic	metamorphic	sedimentary
<input type="checkbox"/>	C	sedimentary	metamorphic	sedimentary
<input type="checkbox"/>	D	metamorphic	sedimentary	metamorphic

(b) The diagram shows reactions of some calcium compounds.



(i) Both calcium carbonate and calcium oxide are white solids.

Suggest how you could show that, when calcium carbonate is heated, a reaction takes place.

(2)

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(ii) Describe what you would observe when water is added, one drop at a time, to cold calcium oxide (**step B**).

(2)

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(iii) Write the balanced equation for the reaction of calcium hydroxide with carbon dioxide (**step C**).

(2)

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(iv) State the common name for calcium hydroxide solution.

(1)

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(c) Explain why calcium carbonate removes acidic gases from emissions in power station chimneys.

(2)

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(Total for Question 4 = 10 marks)



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Electrolysis

5 Electrolysis occurs when solutions of some compounds are decomposed by passing direct electric current through them.

- (a) Sodium chloride solution was electrolysed.
The reaction produced chlorine and hydrogen.
The remaining solution contained sodium hydroxide, NaOH.

(i) State a hazard associated with chlorine gas. (1)

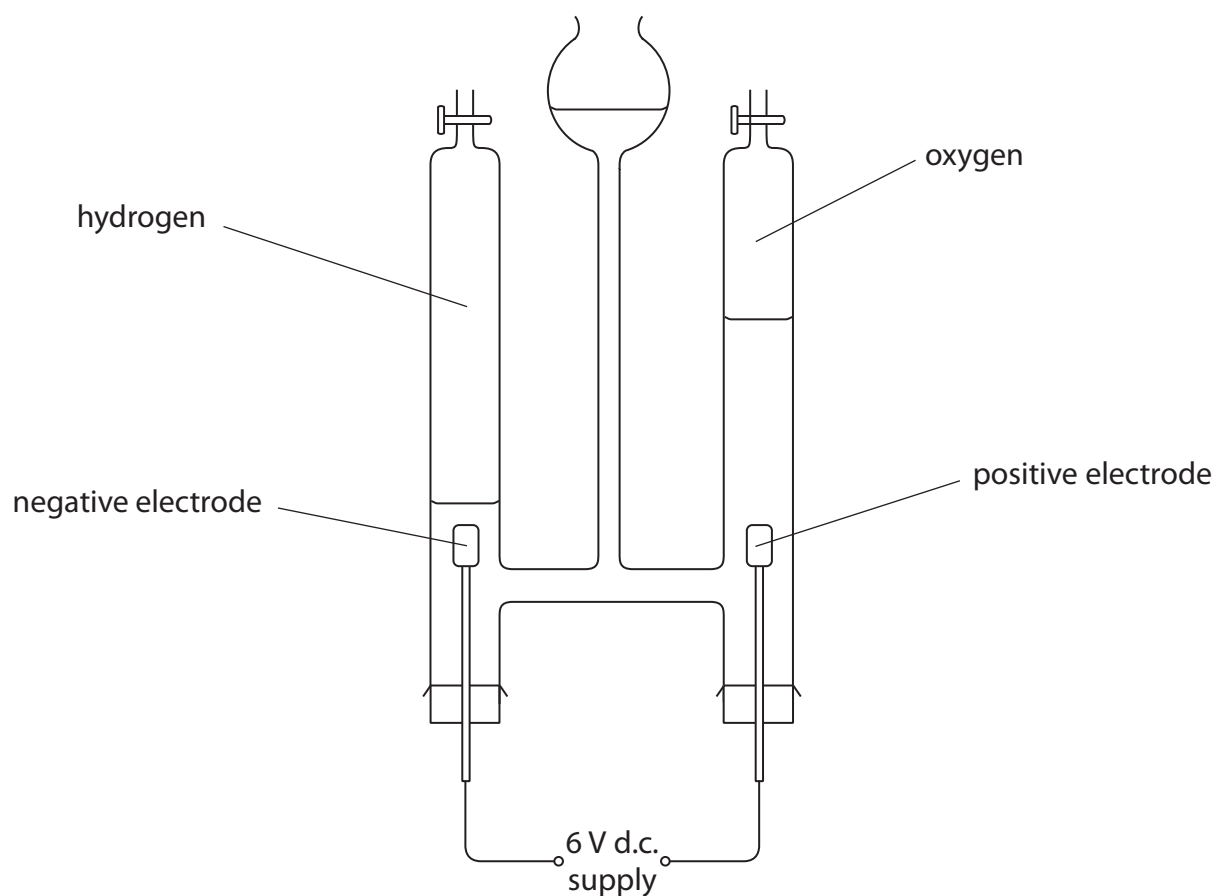
(ii) Describe a test that can be used to identify a sample of gas as chlorine. (2)

(iii) State a use of chlorine. (1)

(iv) Complete and balance the equation for the overall reaction taking place when sodium chloride solution is electrolysed. (2)



*(b) Water was decomposed by electrolysis in the apparatus shown.



The water decomposed into hydrogen and oxygen.
After five minutes, the volumes of hydrogen and oxygen in the tubes were measured.

Two further experiments were carried out changing only one factor in each experiment. All other factors were kept the same.

The table shows the conditions and results of all three experiments.

experiment	time / minutes	current / amps	volume of hydrogen / cm ³	volume of oxygen / cm ³
1	5	0.50	20.0	10.0
2	10	0.50	40.0	20.0
3	5	0.75	30.0	15.0



(c) The photograph shows one of the new buses for London.



The bus uses hydrogen as a fuel.

There are six hydrogen fuel tanks, which can be seen on the roof of the bus. The hydrogen from the tanks reacts with oxygen from the air in a fuel cell to release energy to power the bus.

- (i) Write the balanced equation for the overall reaction that takes place when the hydrogen reacts with oxygen in the fuel cell.

(3)

- (ii) Like all fuels, hydrogen, when mixed with air and ignited, explodes.

Apart from the possibility of an explosion, state another disadvantage of using hydrogen, rather than diesel, as a fuel for buses.

(1)

(Total for Question 6 = 12 marks)

TOTAL FOR PAPER = 60 MARKS



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