



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

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CHEMISTRY

0620/01

Paper 1 Multiple Choice (Core)

For Examination from 2016

SPECIMEN PAPER

45 minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 18.

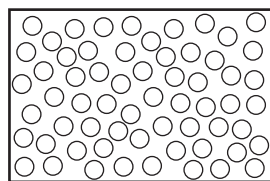
Electronic calculators may be used.

The syllabus is accredited for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

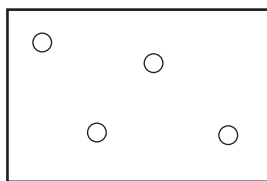
This document consists of **18** printed pages.

2

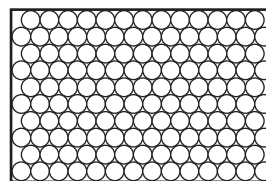
- 1 The diagrams show the arrangement of particles in three different physical states of substance X.



state 1



state 2



state 3

Which statement about the physical states of substance X is correct?

- A** Particles in state 1 vibrate about fixed positions.
B State 1 changes to state 2 by diffusion.
C State 2 changes directly to state 3 by condensation.
D The substance in state 3 has a fixed volume.

- 2 What is always true for a pure substance?

- A** It always boils at 100 °C.
B It contains only one type of atom.
C It has a sharp melting point.
D It is solid at room temperature.

- 3 Element Y has a nucleon number of 19 and a proton number of 9.

Which group in the Periodic Table does it belong to?

- A** I **B** III **C** VII **D** VIII

- 4 The nucleon number and proton number of the lithium atom are shown by the symbol ${}^7_3\text{Li}$.

What is the correct symbol for the lithium ion in lithium chloride?

- A** ${}^6_2\text{Li}^-$ **B** ${}^6_3\text{Li}^+$ **C** ${}^7_3\text{Li}^+$ **D** ${}^7_3\text{Li}^-$

- 5 What is the relative molecular mass, M_r , of HNO_3 ?

- A** 5 **B** 31 **C** 32 **D** 63

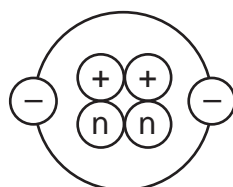
6 The table shows the structure of different atoms and ions.

particle	proton number	nucleon number	number of protons	number of neutrons	number of electrons
Mg	12	24	12	W	12
Mg ²⁺	X	24	12	12	10
F	9	19	9	Y	9
F ⁻	9	19	9	10	Z

What are the values of W, X, Y and Z?

	W	X	Y	Z
A	10	10	9	9
B	10	12	10	9
C	12	10	9	10
D	12	12	10	10

7 The diagram shows the structure of an atom.



key

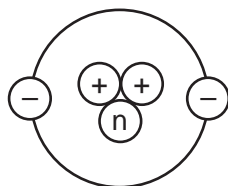
⊕ = proton

⊙ = neutron

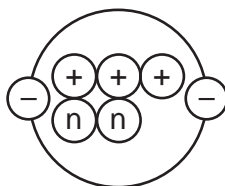
⊖ = electron

Which diagram shows the structure of an isotope of this atom?

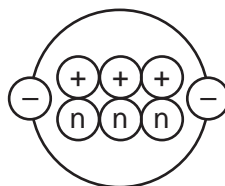
A



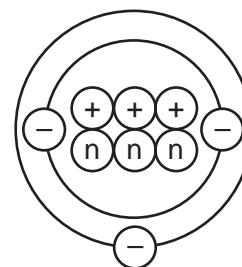
B



C



D



8 Which two elements react together to form an ionic compound?

element	electronic structure
R	2,4
T	2,8
X	2,8,1
Z	2,8,7

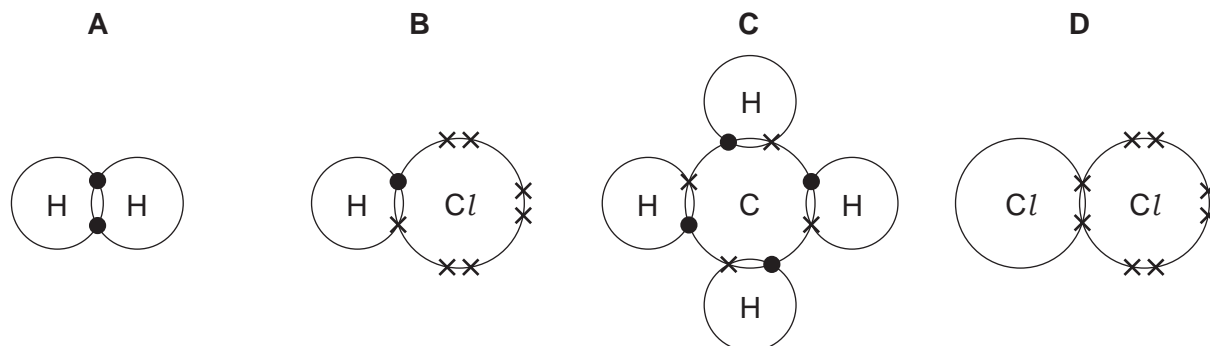
- A** R and T **B** T and X **C** X and Z **D** Z and R

9 Element X forms an acidic, covalent oxide.

Which row shows how many electrons there could be in the outer shell of an atom of X?

	1	2	6	7
A	✓	✓	✗	✗
B	✓	✗	✓	✗
C	✗	✗	✓	✓
D	✗	✓	✗	✓

10 Which diagram does **not** show the outer shell electrons in the molecule correctly?



11 The chemical formulae of two substances, W and X, are given.

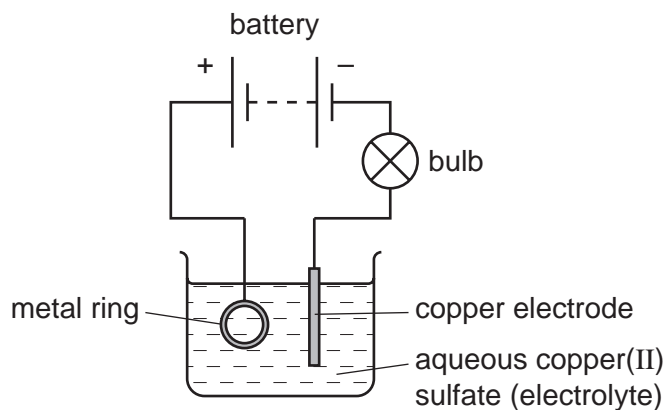


Which statements are correct?

- 1 W and X contain the same amount of oxygen.
- 2 W contains three times as much silicon as X.
- 3 X contains twice as much aluminium as W.

- A** 1 and 2 **B** 1 and 3 **C** 2 and 3 **D** 1, 2 and 3

12 The diagram shows apparatus used in an attempt to electroplate a metal ring with copper.



The experiment did not work.

Which change is needed in the experiment to make it work?

- A Add solid copper(II) sulfate to the electrolyte.
- B Increase the temperature of the electrolyte.
- C Replace the copper electrode with a carbon electrode.
- D Reverse the connections to the battery.

13 Three electrolysis cells are set up. Each cell has inert electrodes.

The electrolytes are listed below.

cell 1 aqueous sodium chloride

cell 2 dilute sulfuric acid

cell 3 molten lead(II) bromide

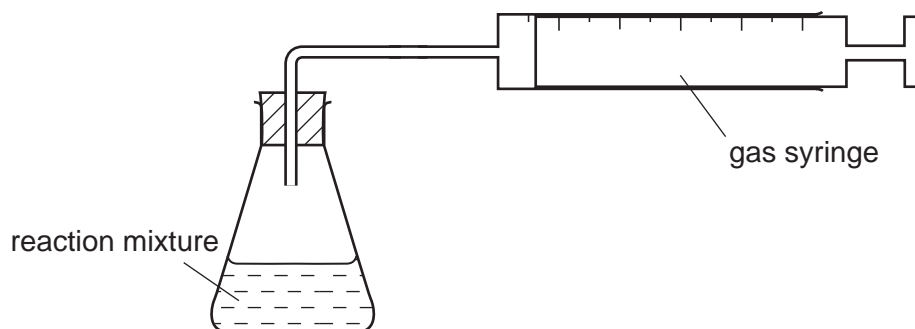
In which of these cells is a gas formed at **both** electrodes?

- A 1 and 2
- B 1 and 3
- C 2 only
- D 3 only

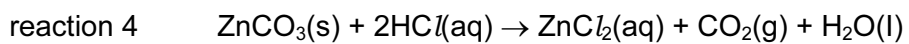
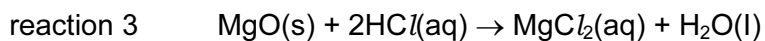
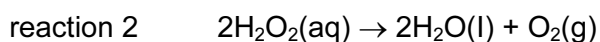
14 Which process is **not** exothermic?

- A burning a fossil fuel
- B obtaining lime from limestone
- C radioactive decay of ^{235}U
- D reacting hydrogen with oxygen

15 The apparatus shown can be used to measure the rate of some chemical reactions.



For which two reactions would this apparatus be suitable?

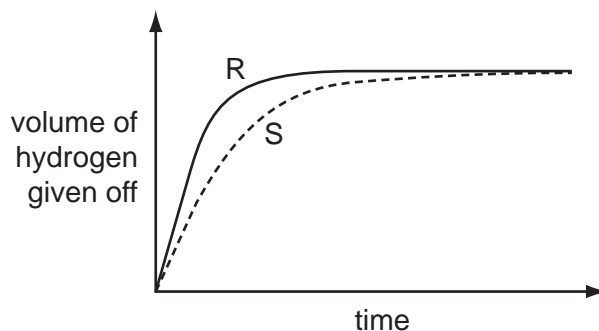


- A** 1 and 2 **B** 1 and 3 **C** 2 and 4 **D** 3 and 4

16 A student investigates the rate of reaction between magnesium and excess sulfuric acid.

The volume of hydrogen given off in the reaction is measured over time.

The graph shows the results of two experiments, R and S.



Which change in conditions would cause the difference between R and S?

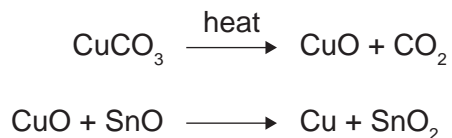
- A A catalyst is added in S.
 - B The acid is more concentrated in R than in S.
 - C The magnesium is less finely powdered in R than in S.
 - D The temperature in R is lower than in S.
- 17 When pink cobalt(II) chloride crystals are heated they form steam and a blue solid.

When water is added to the blue solid, it turns pink and becomes hot.

Which terms describe the pink cobalt(II) chloride crystals and the reactions?

	pink cobalt(II) chloride	reactions
A	aqueous	irreversible
B	anhydrous	reversible
C	hydrated	irreversible
D	hydrated	reversible

- 18 The red colour in some pottery glazes may be formed as a result of the reactions shown.



These equations show that1..... is oxidised and2..... is reduced.

Which substances correctly complete gaps 1 and 2 in the above sentence?

	1	2
A	CO ₂	SnO ₂
B	CuCO ₃	CuO
C	CuO	SnO
D	SnO	CuO

- 19 Carbon dioxide gas reacts with aqueous sodium hydroxide.

Which type of reaction takes place?

- A** decomposition
- B** fermentation
- C** neutralisation
- D** oxidation

- 20 An aqueous solution of the organic compound methylamine has a pH greater than 7.

Which statement about methylamine is correct?

- A** It neutralises an aqueous solution of sodium hydroxide.
- B** It reacts with copper(II) carbonate to give carbon dioxide.
- C** It reacts with hydrochloric acid to form a salt.
- D** It turns blue litmus red.

- 21 A solution contains barium ions and silver ions and one type of anion.

What could the anion be?

- A** chloride only
- B** nitrate only
- C** sulfate only
- D** chloride or nitrate or sulfate

22 A mixture containing two anions was tested and the results are shown below.

test	result
dilute nitric acid added	effervescence of a gas which turned limewater milky
dilute nitric acid added, followed by aqueous silver nitrate	yellow precipitate formed

Which anions were present?

- A carbonate and chloride
- B carbonate and iodide
- C sulfate and chloride
- D sulfate and iodide

23 Astatine is an element in Group VII of the Periodic Table. It has only ever been produced in very small amounts.

What are the likely properties of astatine?

	colour	state	reaction with aqueous potassium iodide
A	black	solid	no reaction
B	dark brown	gas	brown colour
C	green	solid	no reaction
D	yellow	liquid	brown colour

24 The diagram shows the positions of elements L, M, Q, R and T in the Periodic Table.

These letters are not the chemical symbols of the elements.

																		R	T	
L																				
M		Q																		

Which statement about the properties of these elements is correct?

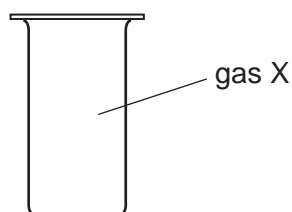
- A L reacts more vigorously with water than does M.
- B L, M and Q are all metals.
- C T exists as diatomic molecules.
- D T is more reactive than R.

25 The table compares the properties of Group I elements with those of transition elements.

Which entry in the table is correct?

	property	Group I elements	transition elements
A	catalytic activity	low	high
B	density	high	low
C	electrical conductivity	low	low
D	melting point	high	low

26 X is a monatomic gas.



Which statement about gas X is correct?

- A X burns in air.
- B X is coloured.
- C X is unreactive.
- D X will displace iodine from potassium iodide.

27 Aluminium is an important metal with many uses.

Some of its properties are listed.

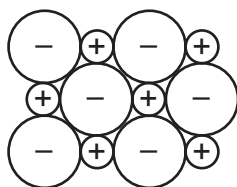
- 1 It is a good conductor of heat.
- 2 It has a low density.
- 3 It has an oxide layer that prevents corrosion.

Which set of properties help to explain the use of aluminium for cooking and storing food?

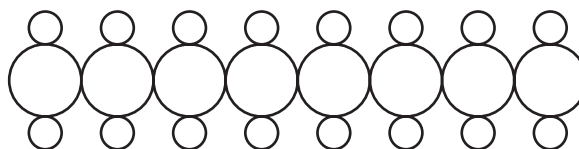
- A** 1 only **B** 1 and 2 only **C** 2 and 3 only **D** 1, 2 and 3

28 Which diagram could represent the structure of an alloy?

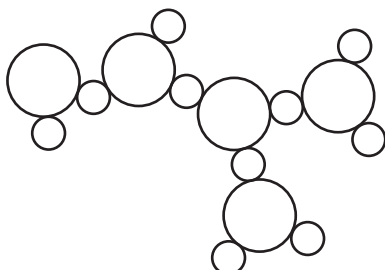
A



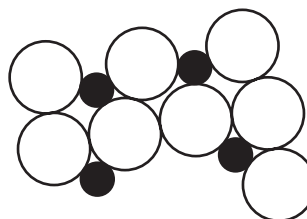
B



C



D



- 29 The table shows the results of adding three metals, P, Q and R, to dilute hydrochloric acid and to water.

metal	dilute hydrochloric acid	water
P	hydrogen produced	hydrogen produced
Q	no reaction	no reaction
R	hydrogen produced	no reaction

What is the order of reactivity of the metals?

	most reactive	→	least reactive
A	P		Q
B	P		R
C	R		P
D	R		Q

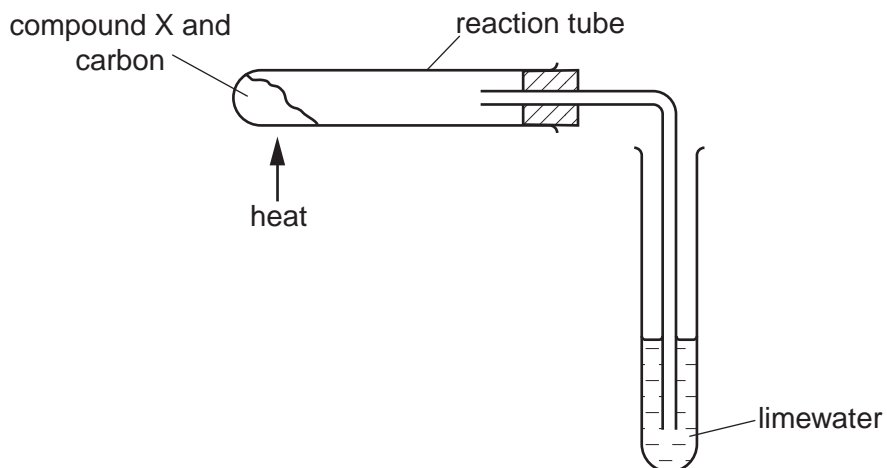
- 30 The table gives the composition of the atmosphere of four newly discovered planets.

planet	composition of atmosphere
W	argon, carbon dioxide and oxygen
X	argon, nitrogen and oxygen
Y	argon, carbon dioxide and methane
Z	methane, nitrogen and oxygen

On which planets is the greenhouse effect likely to occur?

- A** W only
- B** W, X and Z
- C** W and Y only
- D** W, Y and Z

31 Compound X is heated with carbon using the apparatus shown.

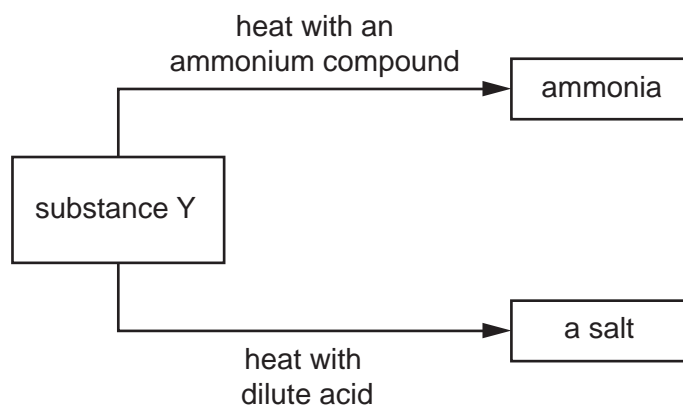


A brown solid is formed in the reaction tube and the limewater turns cloudy.

What is compound X?

- A calcium oxide
- B copper(II) oxide
- C magnesium oxide
- D sodium oxide

32 The diagram shows some reactions of substance Y.

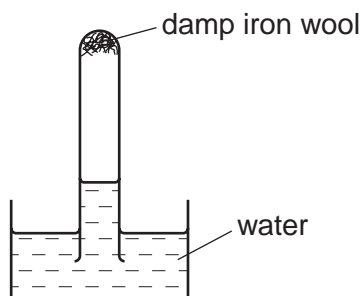


Which type of substance is Y?

- A an alcohol
- B a base
- C a catalyst
- D a metal

33 A test-tube containing damp iron wool is inverted in water.

After three days, the water level inside the test-tube has risen.



Which statement explains this rise?

- A Iron oxide has been formed.
- B Iron wool has been reduced.
- C Oxygen has been formed.
- D The temperature of the water has risen.

34 Greenhouse gases may contribute to climate change.

Two of these gases are emitted into the atmosphere as a result of processes within animals.

Gas1..... is produced by process3..... .

Gas2..... is produced by process4..... .

Which row correctly complete gaps 1, 2, 3 and 4?

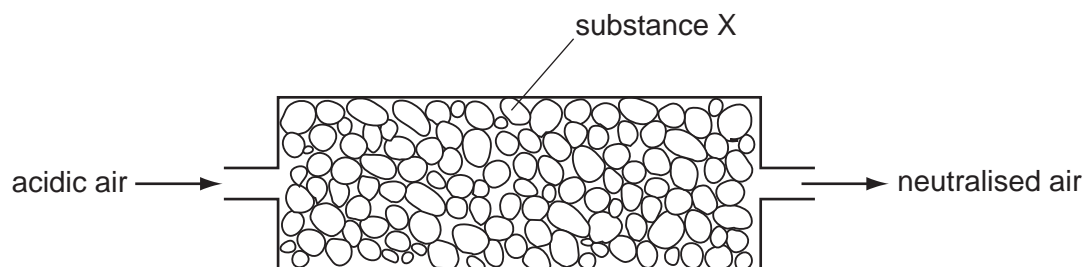
	1	2	3	4
A	CO	C ₂ H ₆	digestion	respiration
B	CO	C ₂ H ₆	respiration	digestion
C	CO ₂	CH ₄	digestion	respiration
D	CO ₂	CH ₄	respiration	digestion

- 35 To grow rose plants, a fertiliser containing nitrogen, phosphorus and potassium is often used. For the best rose flowers, the fertiliser should contain a high proportion of potassium.

Which fertiliser is best for producing rose flowers?

fertiliser	proportion by mass		
	N	P	K
A	9	0	25
B	13	13	20
C	29	5	0
D	29	15	5

- 36 Air containing an acidic impurity was neutralised by passing it through a column containing substance X.

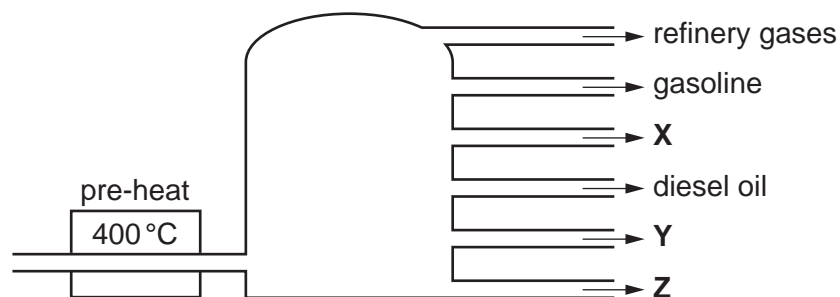


What is substance X?

- A** calcium oxide
- B** sand
- C** sodium chloride
- D** concentrated sulfuric acid

37 In an oil refinery, petroleum is separated into useful fractions.

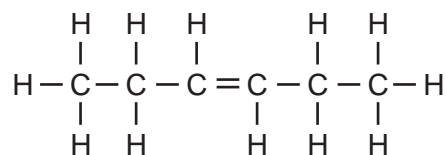
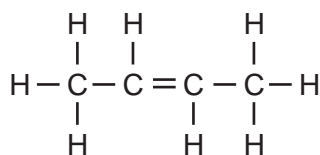
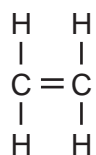
The diagram shows some of these fractions.



What are fractions X, Y and Z?

	X	Y	Z
A	fuel oil	bitumen	paraffin (kerosene)
B	fuel oil	paraffin (kerosene)	bitumen
C	paraffin (kerosene)	bitumen	fuel oil
D	paraffin (kerosene)	fuel oil	bitumen

38 The structures of three compounds are shown.



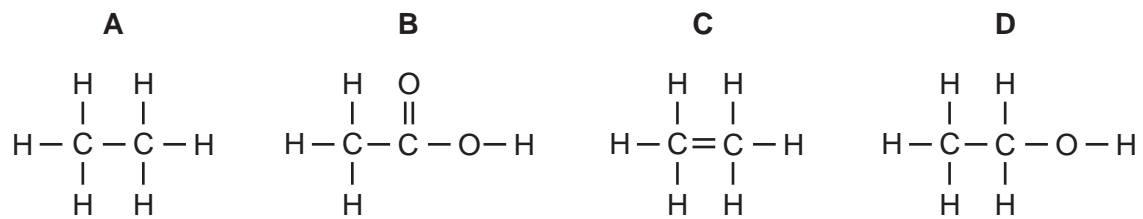
Why do these substances all belong to the same homologous series?

- A** They all contain an even number of carbon atoms.
- B** They all contain the same functional group.
- C** They are all hydrocarbons.
- D** They are all saturated.

39 Which bond is **not** in a molecule of ethanoic acid?

- A** C–O
- B** C=O
- C** C=C
- D** O–H

40 Which structure is **incorrect**?



Group																																	
I	II																III	IV	V	VI	VII	VIII											
	<div style="border: 1px solid black; padding: 5px; width: fit-content; margin: 0 auto;"> Key atomic number atomic symbol name relative atomic mass </div>																																
3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20																
Li lithium 7	Be beryllium 9	H hydrogen 1	B boron 11	C carbon 12	N nitrogen 14	O oxygen 16	F fluorine 19	Ne neon 20	Na sodium 23	Mg magnesium 24	Al aluminium 27	Si silicon 28	P phosphorus 31	S sulfur 32	Cl chlorine 35.5	Ar argon 40	He helium 4																
19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36																
K potassium 39	Ca calcium 40	Sc scandium 45	Ti titanium 48	V vanadium 51	Cr chromium 52	Mn manganese 55	Fe iron 56	Co cobalt 59	Ni nickel 59	Cu copper 64	Zn zinc 65	Ga gallium 70	Ge germanium 73	As arsenic 75	Se selenium 79	Br bromine 80	Kr krypton 84																
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54																
Rb rubidium 85	Sr strontium 88	Y yttrium 89	Zr zirconium 91	Nb niobium 93	Mo molybdenum 96	Tc technetium —	Ru ruthenium 101	Rh rhodium 103	Pd palladium 106	Ag silver 108	Cd cadmium 112	In indium 115	Sn tin 117	Sb antimony 122	Te tellurium 128	I iodine 127	Xe xenon 131																
55	56	57–71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86																
Cs caesium 133	Ba barium 137	lanthanoids	Hf hafnium 178	Ta tantalum 181	W tungsten 184	Re rhenium 186	Os osmium 190	Ir iridium 192	Pt platinum 195	Au gold 197	Hg mercury 201	Tl thallium 204	Pb lead 207	Bi bismuth 209	Po polonium —	At astatine —	Rn radon —																
87	88	89–103	104	105	106	107	108	109	110	111	112	114	114	116	116	—	—																
Fr francium —	Ra radium —	actinoids	Rf rutherfordium —	Db dubnium —	Sg seaborgium —	Bh bohrium —	Hs hassium —	Mt meitnerium —	Ds darmstadtium —	Rg roentgenium —	Cn copernicium —	Fl flerovium —	Fl flerovium —	Lv livermorium —	—	—	—																
57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	71	71	71																
La lanthanum 139	Ce cerium 140	Pr praseodymium 141	Nd neodymium 144	Pm promethium —	Sm samarium 150	Eu europium 152	Gd gadolinium 157	Tb terbium 159	Dy dysprosium 163	Ho holmium 165	Er erbium 167	Tm thulium 169	Yb ytterbium 173	Lu lutetium 175	Lu lutetium 175	Lu lutetium 175	Lu lutetium 175																
89	90	91	92	93	94	95	96	97	98	99	100	101	102	103	103	103	103																
Ac actinium —	Th thorium 232	Pa protactinium 231	U uranium 238	Np neptunium —	Pu plutonium —	Am americium —	Cm curium —	Bk berkelium —	Cf californium —	Es einsteinium —	Fm fermium —	Md mendelevium —	No nobelium —	Lr lawrencium —	Lr lawrencium —	Lr lawrencium —	Lr lawrencium —																

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.)

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