



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CANDIDATE
NAME

CENTRE
NUMBER

--	--	--	--	--

CANDIDATE
NUMBER

--	--	--	--

CHEMISTRY

0620/23

Paper 2

October/November 2014

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name in the spaces at the top of this page.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 16.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

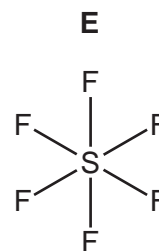
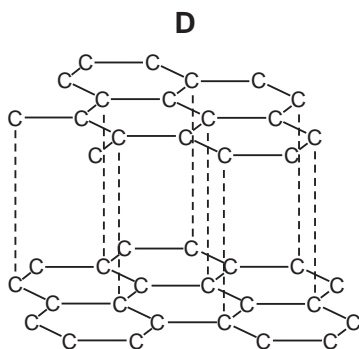
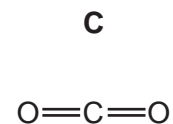
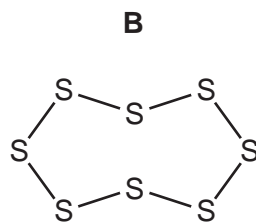
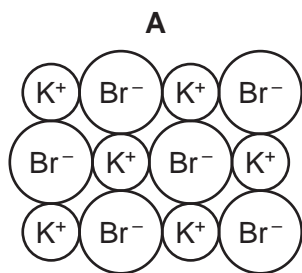
The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **16** printed pages.



2

1 (a) The structures of five substances, **A**, **B**, **C**, **D** and **E**, are shown below.



Answer the following questions about these substances.
Each substance may be used once, more than once or not at all.

- (i) Which **two** substances are elements? and [2]
- (ii) Which substance has a giant ionic structure? [1]
- (iii) Which substance is a product formed when a hydrocarbon is completely burnt in air?
..... [1]
- (iv) Which substance, when molten, produces a brown vapour at the anode when electrolysed?
..... [1]
- (v) Which substance is used as a lubricant? [1]
- (b) Deduce the simplest formula of substance **A**.
..... [1]
- (c) Calculate the relative molecular mass of substance **E**.
You must show your working.

[2]

[Total: 9]

2 (a) Use ideas about the movement and arrangement of particles to explain why:

- solids have a definite volume and shape,
- liquids have a definite volume but no definite shape,
- gases have no definite volume or shape.

.....

.....

.....

.....

.....

..... [4]

(b) The table shows some properties of six substances, **A** to **F**, which are either solids or liquids at room temperature.

substance	melting point /°C	electrical conductivity	solubility in water
A	+3550	does not conduct in any state	insoluble
B	+44	does not conduct in any state	insoluble
C	+1660	conducts when solid or liquid	insoluble
D	+681	only conducts when in aqueous solution or liquid	soluble
E	-39	conducts when solid or liquid	insoluble
F	-11	does not conduct in any state	insoluble

(i) Which substance has a giant covalent structure?

..... [1]

(ii) Which **two** substances are liquids at room temperature?

..... [2]

(iii) Which **two** substances are metals?

..... [2]

(iv) Which **two** substances are simple molecules?

..... [2]

4

(c) Dry air contains mainly nitrogen, noble gases and oxygen.

(i) Which **one** of the following shows the correct composition of dry air?
Tick **one** box.

nitrogen 21%, oxygen 78%, noble gases 1%

nitrogen 1%, oxygen 78%, noble gases 21%

nitrogen 69%, oxygen 21%, noble gases 10%

nitrogen 78%, oxygen 21%, noble gases 1%

[1]

(ii) Metals can be joined together by high temperature welding.
This process is sometimes carried out in the presence of argon.
Suggest why welding is carried out in the presence of argon.

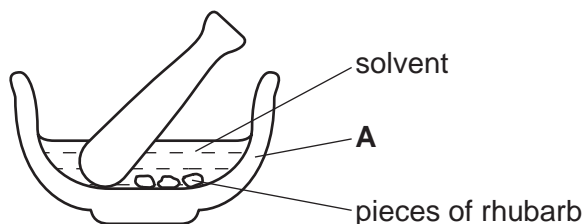
.....

..... [2]

[Total: 14]

3 Rhubarb is a plant which has a red stem.

(a) A student separated the pigments in the rhubarb stem by chromatography. He used the apparatus shown below to extract the pigments.



(i) State the name of the piece of apparatus labelled **A**.

..... [1]

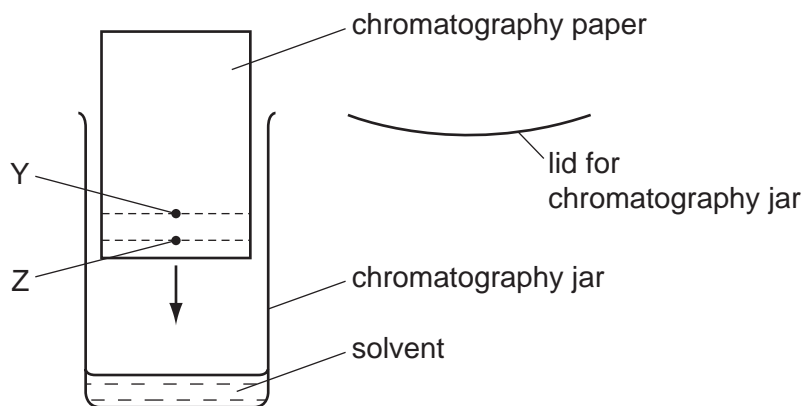
(ii) Suggest a suitable solvent, other than water, that he could use to extract the pigments.

..... [1]

(iii) The solution of pigments was not concentrated enough to use for chromatography. Suggest how the student could make the solution more concentrated.

..... [1]

(b) The student carried out chromatography using the apparatus shown below.



(i) A spot of the pigment mixture was placed at Y. Explain why a spot of the mixture was not placed at Z.

..... [1]

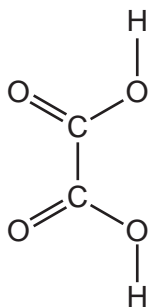
(ii) Describe how the rest of the procedure was carried out.

.....

 [2]

6

- (c) Rhubarb leaves contain ethanedioic acid.
The structure of ethanedioic acid is shown below.

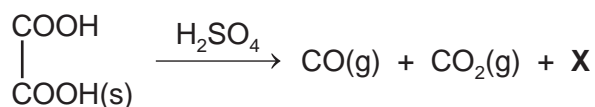


- (i) On the structure above, put a ring around a carboxylic acid group. [1]

- (ii) Deduce the molecular formula of ethanedioic acid.

..... [1]

- (d) A teacher heated ethanedioic acid with concentrated sulfuric acid.
The equation for the reaction is:



- (i) Deduce the formula of compound X.

..... [1]

- (ii) At the end of the reaction, the contents of the test-tube contained diluted sulfuric acid only.
Explain why.

..... [1]

- (iii) Carbon dioxide is a product of this reaction.
State **one** common source of the carbon dioxide in the atmosphere.

..... [1]

- (iv) Explain why an increase in the concentration of carbon dioxide in the atmosphere is harmful to the environment.

.....

..... [2]

[Total: 13]

4 A mixture of soil and water was shaken and then filtered.

- (a) Draw a labelled diagram of the apparatus you would use for separating the insoluble particles of soil from the solution.

[2]

- (b) The filtrate was then evaporated.

The table shows the composition and mass of each compound obtained by evaporating the filtrate.

compound	ions present in the compound	mass of compound/g
calcium carbonate	Ca^{2+} and CO_3^{2-}	4.0
calcium sulfate	Ca^{2+} and SO_4^{2-}	5.0
magnesium sulfate	Mg^{2+} and SO_4^{2-}	2.8
	K^+ and NO_3^-	1.2
potassium sulfate	K^+ and SO_4^{2-}	2.4
sodium carbonate		3.0
sodium chloride	Na^+ and Cl^-	1.6

- (i) State the name of the compound which contains K^+ and NO_3^- ions.

..... [1]

- (ii) Write the symbols for the ions present in sodium carbonate.

..... [1]

- (iii) Which compound with a singly charged negative ion has the highest mass in the mixture?

..... [1]

The table from page 7 is repeated below:

compound	ions present in the compound	mass of compound/g
calcium carbonate	Ca^{2+} and CO_3^{2-}	4.0
calcium sulfate	Ca^{2+} and SO_4^{2-}	5.0
magnesium sulfate	Mg^{2+} and SO_4^{2-}	2.8
	K^+ and NO_3^-	1.2
potassium sulfate	K^+ and SO_4^{2-}	2.4
sodium carbonate		3.0
sodium chloride	Na^+ and Cl^-	1.6

(iv) Calculate:

the total mass of all the compounds present in the mixture,

..... [1]

the percentage of magnesium sulfate by mass in the mixture.

[1]

(c) Calcium carbonate decomposes when heated.

(i) Complete the symbol equation for this reaction.



[1]

(ii) Calcium oxide, CaO, reacts with water to form a strongly alkaline solution.

Which **one** of the following pH values is strongly alkaline?

Put a ring around the correct answer.

pH 3

pH 7

pH 8

pH 12

[1]

[Total: 9]

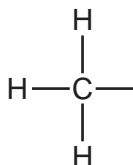
- 5 Methanol reacts with excess hydrochloric acid to form chloromethane and water.



- (a) To which homologous series does methanol belong?

..... [1]

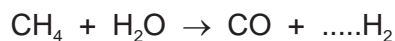
- (b) Complete the structure of methanol below to show its functional group.



[1]

- (c) Methanol can be made from synthesis gas which contains carbon monoxide and hydrogen. Synthesis gas is made from methane.

- (i) Complete the symbol equation for this reaction.

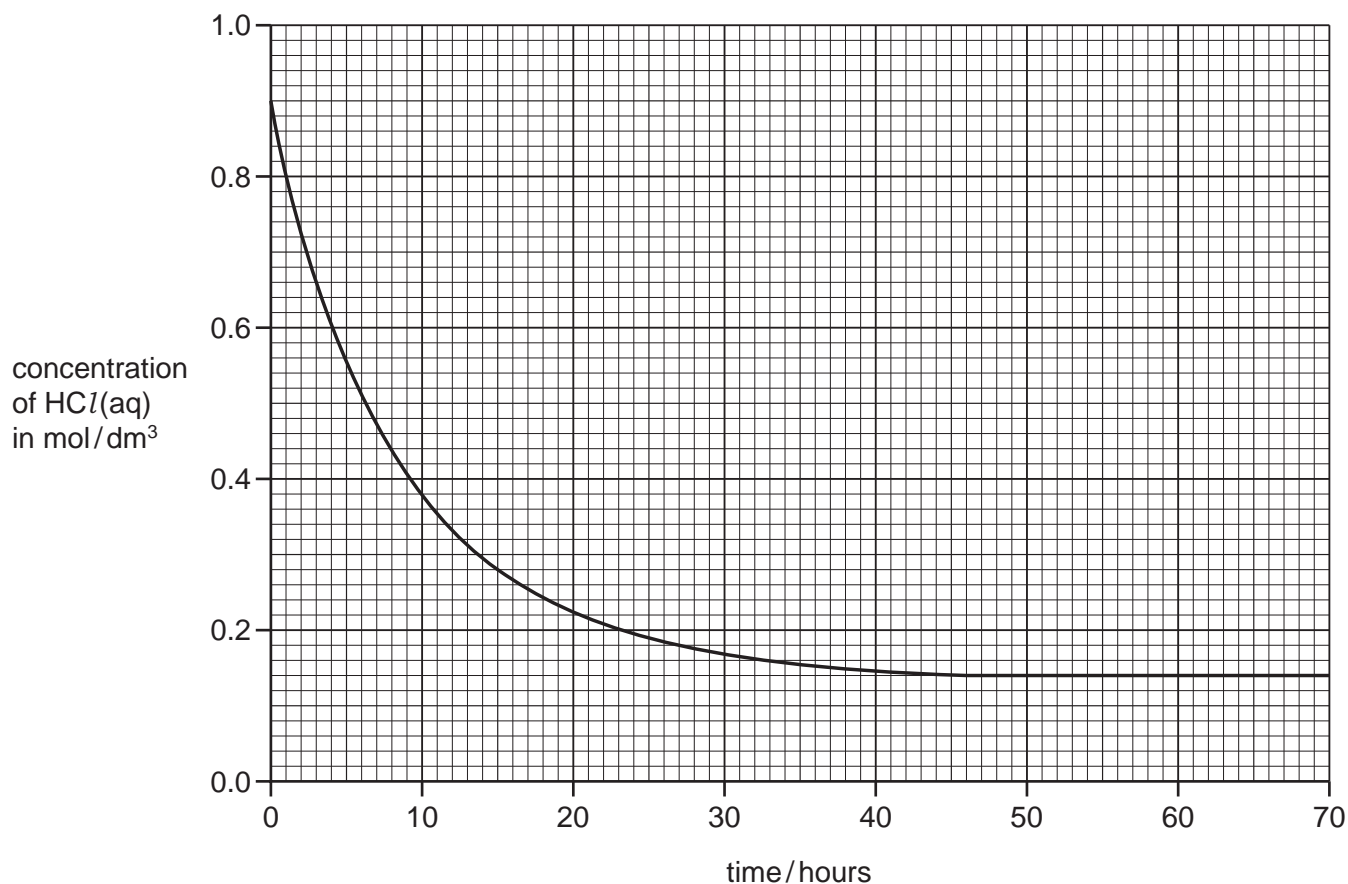


[1]

- (ii) Suggest **two** hazards associated with the products of this reaction.

.....
 [2]

- (d) A student investigated the rate of reaction of methanol with hydrochloric acid. The graph below shows how the concentration of hydrochloric acid changes with time.



- (i) Describe how the concentration of hydrochloric acid changes with time.

.....
 [2]

- (ii) Deduce the concentration of hydrochloric acid when the reaction had proceeded for 15 hours.

..... [1]

- (iii) At what time was the reaction just complete?

..... [1]

- (iv) On the grid above, draw a line to show how the concentration of hydrochloric acid changes with time when the reaction takes place at a higher temperature. [2]

- (e) Hydrochloric acid is formed when hydrogen chloride gas is dissolved in water.
Draw a dot-and-cross diagram to show the electron arrangement in a molecule of hydrogen chloride.
Show only the outer electron shells.
Show hydrogen electrons as x.
Show chlorine electrons as ●.

[2]

[Total: 13]

6 Sulfur burns in air to form sulfur dioxide.

(a) (i) Is sulfur dioxide an acidic or basic oxide?

Give a reason for your answer.

.....
 [1]

(ii) Sulfur dioxide is an atmospheric pollutant.

Explain why sulfur dioxide in the atmosphere can erode buildings made of limestone.

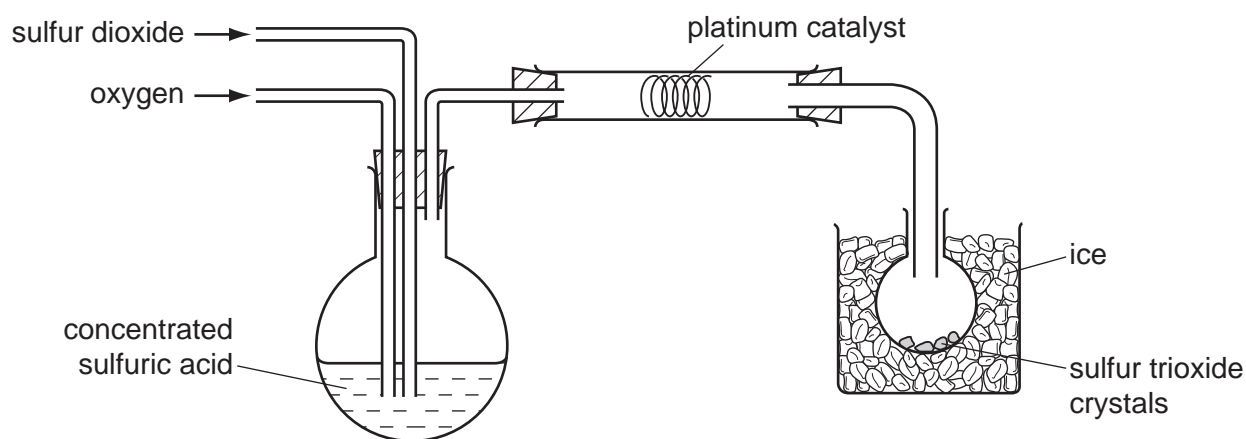
.....

 [3]

(b) Sulfur dioxide reacts with oxygen to form sulfur trioxide.

Sulfur trioxide can be made in the laboratory using the apparatus shown below.

Sulfur trioxide has a melting point of 17 °C and a boiling point of 45 °C.



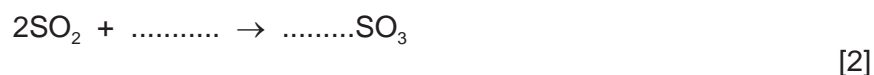
(i) Suggest **one** safety precaution when carrying out this experiment.

..... [1]

(ii) What is the purpose of the platinum catalyst?

..... [1]

(iii) Complete the symbol equation for the reaction.



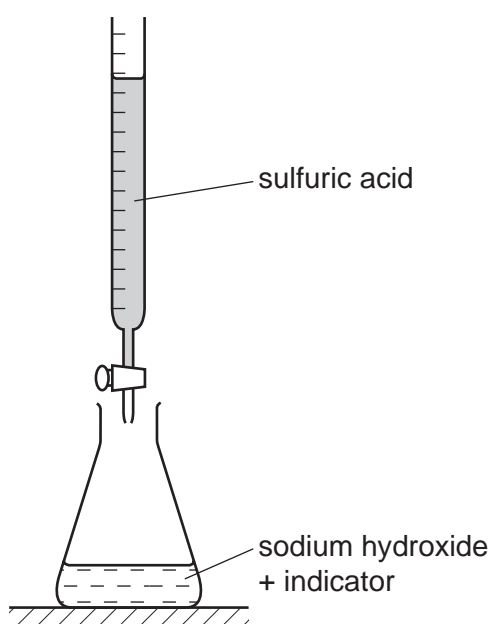
(iv) Suggest why the sulfur trioxide is collected in a flask surrounded by ice.

..... [1]

- (v) When 64 g of sulfur dioxide react with excess oxygen, 80 g of sulfur trioxide is formed. Calculate the mass of sulfur trioxide formed from 160 g of sulfur dioxide.

mass = g [1]

- (c) Sulfur trioxide reacts with water to form sulfuric acid.
A student used the apparatus shown below to determine the concentration of a solution of sodium hydroxide.



- (i) Which **one** of these pieces of apparatus should the student use to put 25.0 cm³ of sodium hydroxide into the flask.
Tick **one** box.

beaker

measuring cylinder

test-tube

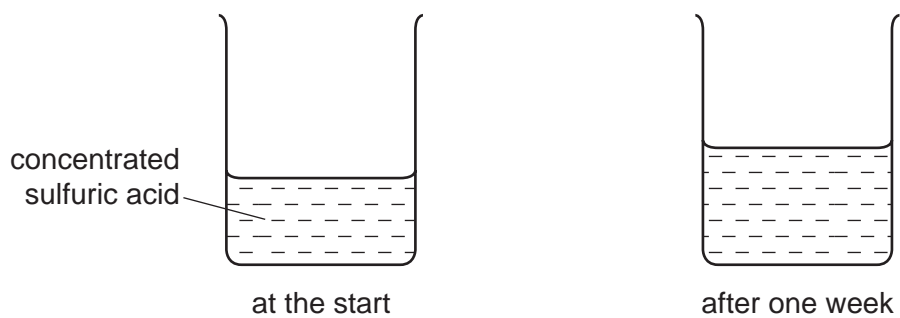
volumetric pipette

[1]

- (ii) How would the student know when the sulfuric acid had neutralised the sodium hydroxide?

..... [1]

- (d) Clean air contains mainly nitrogen, noble gases, oxygen and water vapour. A teacher left a beaker of concentrated sulfuric acid open to the air for a week. After a week, the concentration of sulfuric acid in the beaker had decreased.



Explain these results by referring to one or more of the substances present in the air.

.....
..... [1]

[Total: 13]

7 (a) Describe the properties of chlorine, bromine and iodine.

In your answer, include the trends in:

- their state,
- their colour,
- their reactivity.

.....

.....

.....

.....

..... [4]

(b) A molecule of chlorine can be written as $Cl-Cl$.
Which **one** of the following words describes this molecule?
Put a ring around the correct answer.

diatomic giant ionic monatomic

[1]

(c) Draw the electronic structure of a fluorine atom.

[2]

(d) The equation below describes the reaction of a halogen with a halide.



Write a word equation for this reaction.

[2]

[Total: 9]

DATA SHEET
The Periodic Table of the Elements

		Group																																																																																																																															
I	II	III	IV	V	VI	VII	0																																																																																																																										
7 Li Lithium 3	9 Be Beryllium 4	1 H Hydrogen 1	11 B Boron 5	12 C Carbon 6	13 Al Aluminium 13	14 N Nitrogen 7	15 O Oxygen 8	16 F Fluorine 9	17 Ne Neon 10	18 Ar Argon 18	19 K Potassium 19	20 Ca Calcium 20	21 Sc Scandium 21	22 Ti Titanium 22	23 V Vanadium 23	24 Cr Chromium 24	25 Mn Manganese 25	26 Fe Iron 26	27 Co Cobalt 27	28 Ni Nickel 28	29 Cu Copper 29	30 Zn Zinc 30	31 Ga Gallium 31	32 Ge Germanium 32	33 As Arsenic 33	34 Se Selenium 34	35 Br Bromine 35	36 Kr Krypton 36	37 Rb Rubidium 37	38 Sr Strontium 38	39 Y Yttrium 39	40 Zr Zirconium 40	41 Nb Niobium 41	42 Mo Molybdenum 42	43 Tc Technetium 43	44 Ru Ruthenium 44	45 Rh Rhodium 45	46 Pd Palladium 46	47 Ag Silver 47	48 Cd Cadmium 48	49 In Indium 49	50 Sn Tin 50	51 Sb Antimony 51	52 Te Tellurium 52	53 I Iodine 53	54 Xe Xenon 54	55 Cs Caesium 55	56 Ba Barium 56	57 La Lanthanum 57	72 Hf Hafnium 72	73 Ta Tantalum 73	74 W Tungsten 74	75 Re Rhenium 75	76 Os Osmium 76	77 Ir Iridium 77	78 Pt Platinum 78	79 Au Gold 79	80 Hg Mercury 80	81 Tl Thallium 81	82 Pb Lead 82	83 Bi Bismuth 83	84 Po Polonium 84	85 At Astatine 85	86 Rn Radon 86	87 Fr Francium 87	88 Ra Radium 88	89 Ac Actinium 89	†	90 Th Thorium 90	91 Pa Protactinium 91	92 U Uranium 92	93 Np Neptunium 93	94 Pu Plutonium 94	95 Am Americium 95	96 Cm Curium 96	97 Bk Berkelium 97	98 Cf Californium 98	99 Es Einsteinium 99	100 Fm Fermium 100	101 Md Mendelevium 101	102 No Nobelium 102	103 Lr Lawrencium 103	133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	192 Ir Iridium 77	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	212 Po Polonium 84	214 At Astatine 85	216 Rn Radon 86	226 Ra Radium 88	227 Ac Actinium 89	†	232 Th Thorium 90	238 U Uranium 92	238 Np Neptunium 93	238 Pu Plutonium 94	238 Am Americium 95	238 Cm Curium 96	238 Bk Berkelium 97	238 Cf Californium 98	238 Es Einsteinium 99	238 Fm Fermium 100	238 Md Mendelevium 101	238 No Nobelium 102	238 Lr Lawrencium 103	140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	159 Tb Terbium 65	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71

*58-71 Lanthanoid series
†90-103 Actinoid series

Key

a	X
b	

a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included the publisher will be pleased to make amends at the earliest possible opportunity.

Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.