



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CANDIDATE
NAME

CENTRE
NUMBER

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CHEMISTRY

0620/22

Paper 2

October/November 2014

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name in the spaces at the top of this page.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 20.

You may lose marks if you do not show your working or if you do not use appropriate units.

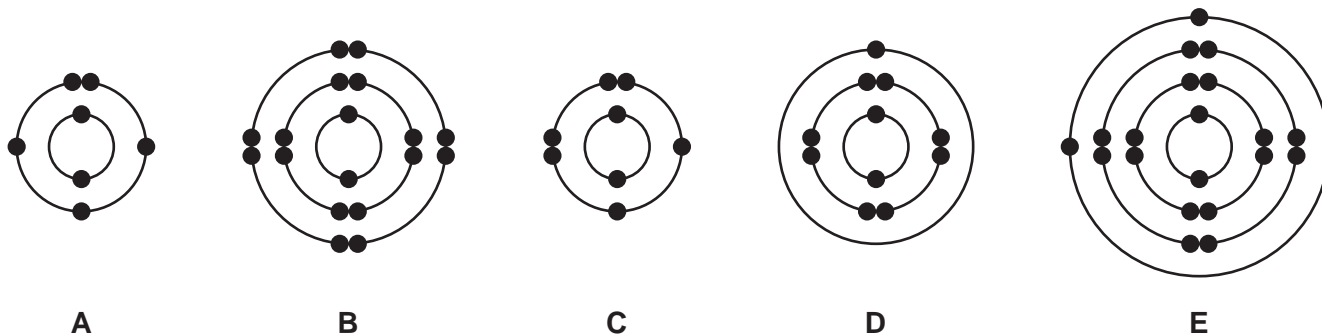
At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **17** printed pages and **3** blank pages.

1 (a) The electronic structure of five atoms, **A**, **B**, **C**, **D** and **E**, are shown below.



Answer the following questions about these structures.
Each structure can be used once, more than once or not at all.

Which structure:

- (i) represents an atom of an element in Group V of the Periodic Table, [1]
- (ii) has a complete outer shell of electrons, [1]
- (iii) represents an oxygen atom, [1]
- (iv) has a proton number of 20, [1]
- (v) is an atom of an element in Period 4 of the Periodic Table, [1]
- (vi) has a single valency electron? [1]

(b) Complete the following sentences about isotopes using words from the list below.

atoms ions molecules neutrons nuclei protons

Isotopes are of the same element with the same number of
but different numbers of [3]

[Total: 9]

2 The table below shows some nutritional information on a bottle of apple juice.

contents	mass present in g/100 cm ³
protein	0.10
sugars	10.40
unsaturated fat	0.10
saturated fat	0.06
chloride ions, Cl ⁻	0.04
magnesium ions, Mg ²⁺	0.01
nitrate ions, NO ₃ ⁻	0.01
potassium ions, K ⁺	0.02
sodium ions, Na ⁺	0.05
X , SO ₄ ²⁻	0.01

(a) Answer these questions using information from the table.

(i) Which negatively charged ion is present in the highest concentration?

..... [1]

(ii) State the name of the ion, **X**, whose formula is SO₄²⁻.

..... [1]

(iii) The formulae for some chlorides are shown below.

aluminium chloride, AlCl₃

calcium chloride, CaCl₂

lead(IV) chloride, PbCl₄

potassium chloride, KCl

Deduce the formula for magnesium chloride.

..... [1]

(iv) Calculate the mass of sugars in 250 cm³ of this apple juice.

..... g [1]

4

- (b) The fats in the apple juice are both saturated and unsaturated.
Describe a test to distinguish between saturated and unsaturated compounds.

test

result with saturated compound

result with unsaturated compound

[3]

- (c) Apple juice is slightly acidic.

- (i) Which **one** of the following pH values is slightly acidic?
Put a ring around the correct answer.

pH 1

pH 5

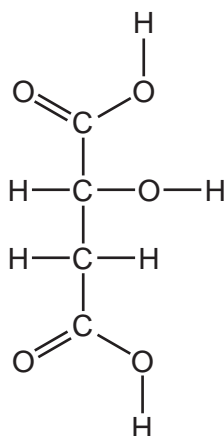
pH 7

pH 9

pH 14

[1]

- (ii) One of the acids found in apple juice is malic acid.
The structure of malic acid is shown below.

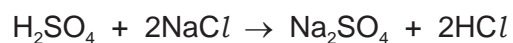


On the structure of malic acid above, put a ring around a carboxylic acid functional group. [1]

[Total: 9]

5

- 3 Hydrogen chloride gas can be prepared by the action of concentrated sulfuric acid on sodium chloride.

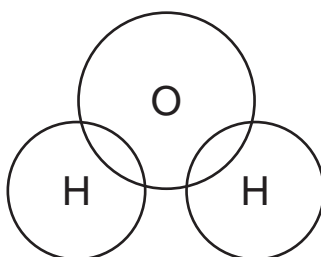


- (a) Write the word equation for this reaction.

..... [1]

- (b) Hydrogen chloride dissolves in water to form hydrochloric acid.

- (i) Complete the dot-and-cross diagram to show the arrangement of the outer shell electrons in water.

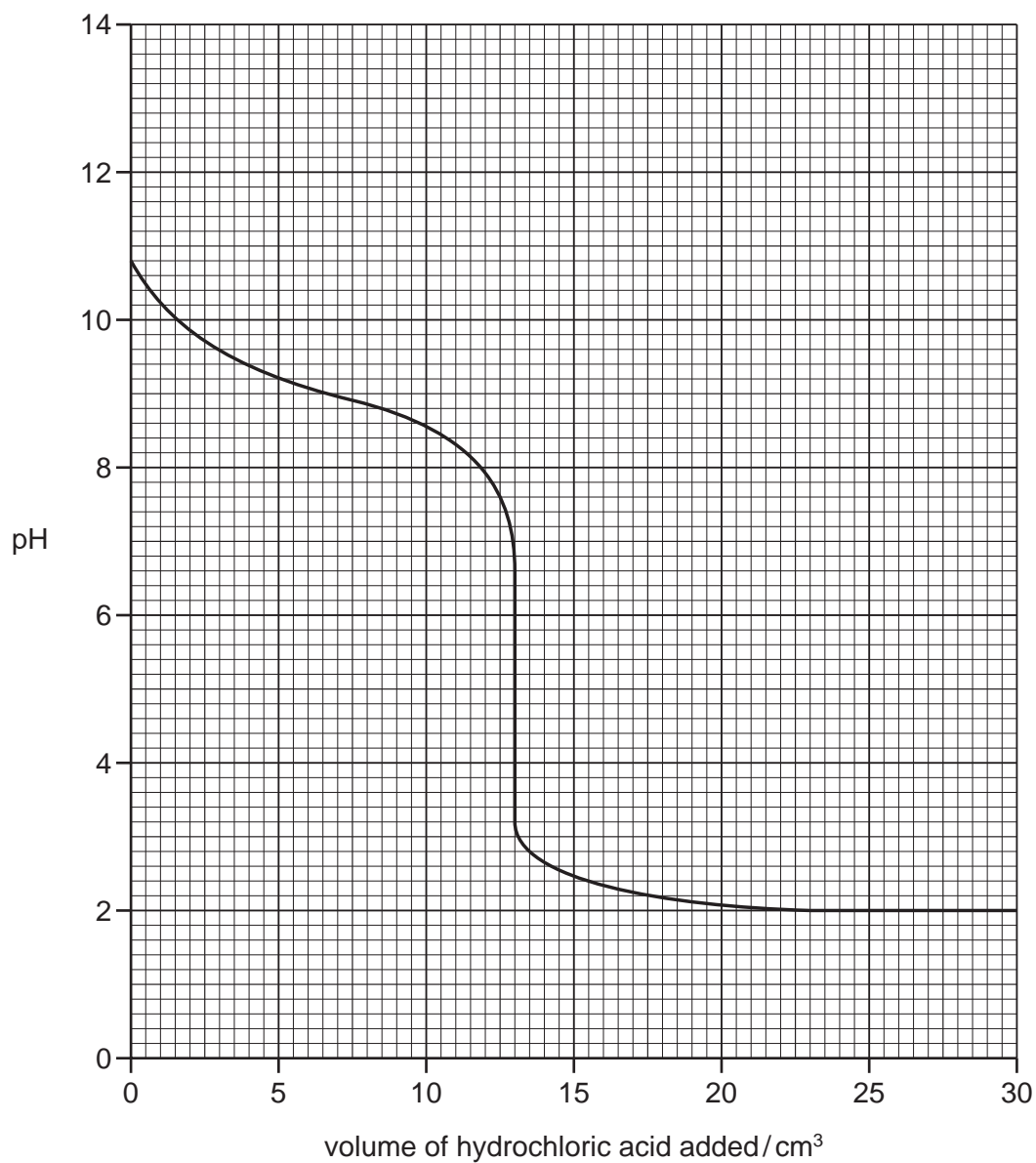


[2]

- (ii) Describe what you would observe when a few drops of silver nitrate solution are added to hydrochloric acid.

..... [2]

- (c) The graph below shows how pH changes when aqueous ammonia is neutralised by hydrochloric acid.



- (i) What is the pH of the aqueous ammonia at the start of the experiment?

..... [1]

- (ii) What volume of hydrochloric acid has been added when the pH is 10?

..... [1]

- (iii) What volume of hydrochloric acid has been added when the pH is changing most quickly?

..... [1]

- (d) Concentrated hydrochloric acid reduces manganese(IV) oxide, MnO_2 , to manganese(II) chloride.



How does this equation show that manganese(IV) oxide gets reduced?

..... [1]

- (e) The table shows some properties of four metals, **A**, **B**, **C** and **D**, and their oxides.

metal	density in g/cm^3	boiling point / $^\circ\text{C}$	colour of oxide	charge on the metal ion
A	2.99	2831	white	3+
B	0.53	1342	white	1+
C	7.86	2750	black or red-brown	2+ or 3+
D	7.14	907	white	2+

Which **one** of these metals is a transition metal?

Use the information in the table to explain your answer.

.....

 [2]

[Total: 11]

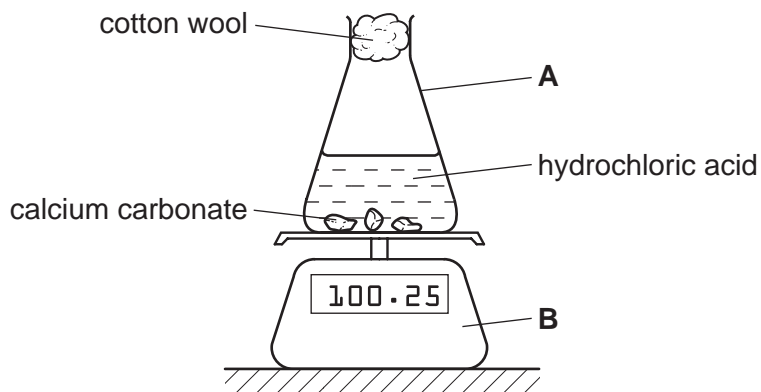
4 Calcium carbonate reacts with dilute hydrochloric acid.

(a) Complete the symbol equation for this reaction.



[2]

(b) The rate of this reaction can be followed using the apparatus shown below.



(i) State the names of the pieces of apparatus labelled **A** and **B**.

A

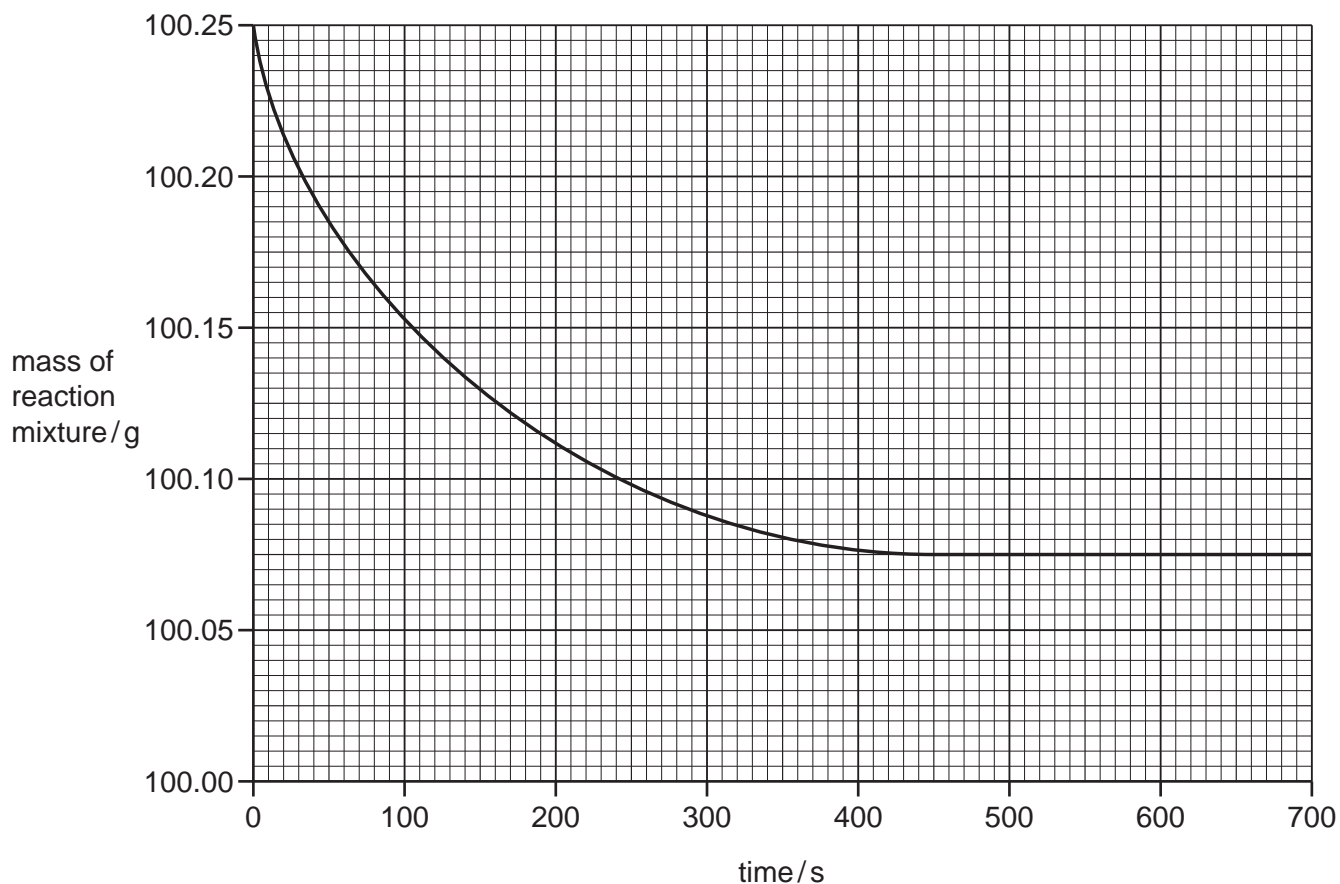
B

[2]

(ii) Explain why the mass of the reaction mixture decreases with time.

..... [1]

- (c) The graph below shows how the mass of the reaction mixture changes with time. The calcium carbonate was in excess and large pieces of calcium carbonate were used.



- (i) At what time was the reaction just complete?

..... [1]

- (ii) Calculate the total loss in mass of the reaction mixture in this experiment.

..... [1]

- (iii) How does the rate of reaction change when:

smaller pieces of calcium carbonate are used,

.....

the temperature is decreased,

.....

the concentration of hydrochloric acid is decreased?

.....

[3]

(d) When heated, calcium carbonate breaks down to form calcium oxide and carbon dioxide.

Which **two** words from the list below describe this reaction?

Tick **two** boxes.

combustion

decomposition

endothermic

exothermic

oxidation

[2]

(e) Calcium oxide is used in flue-gas desulfurisation.

(i) Explain how flue-gas desulfurisation works.

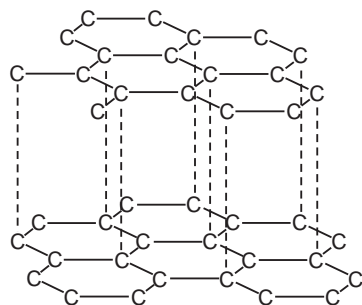
.....
.....
..... [2]

(ii) Give **one** other use of calcium oxide.

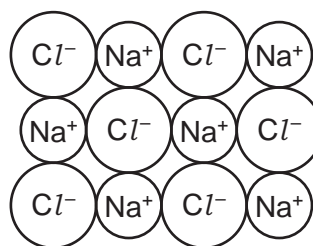
..... [1]

[Total: 15]

5 The structures of graphite and sodium chloride are shown below.



graphite



sodium chloride

(a) Describe the similarity and differences in these structures.

.....

.....

.....

.....

.....

..... [4]

(b) Graphite is a form of carbon.
Carbon is an element.

(i) What is meant by the term *element*?

.....

..... [1]

(ii) Write a symbol equation for the complete combustion of carbon.

[2]

(c) The table shows some properties of four substances, **A**, **B**, **C** and **D**.

substance	melting point /°C	boiling point /°C	electrical conductivity
A	-7	+59	does not conduct
B	-157	-152	does not conduct
C	+769	+1930	conducts when molten but not when solid
D	+1410	+2355	does not conduct

Which **one** of these substances, **A**, **B**, **C** or **D**,

- (i) is a liquid at room temperature, [1]
- (ii) is a giant ionic structure, [1]
- (iii) is a noble gas, [1]
- (iv) is a giant covalent structure? [1]

[Total: 11]

- 6 The table below shows some properties of the first five members of the alkane homologous series.

alkane	molecular formula	boiling point /°C	density of the liquid alkane in g/cm ³
methane	CH ₄	-164	0.47
ethane	C ₂ H ₆	-88	
propane	C ₃ H ₈	-42	0.59
butane	C ₄ H ₁₀	0	0.60
pentane		+36	0.63

- (a) (i) What do you understand by the term *homologous series*?

.....
 [2]

- (ii) Deduce the molecular formula for pentane.

..... [1]

- (iii) Describe how the boiling points of these alkanes change as the number of carbon atoms increases.

..... [1]

- (iv) Deduce the density of liquid ethane.

..... [1]

- (b) Methane is a fuel which is a gas at room temperature.

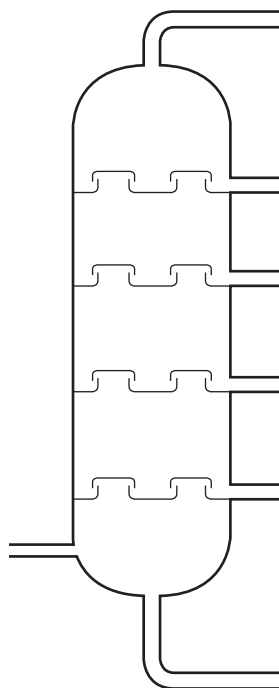
State the name of a fuel which is:

a solid at room temperature,

a liquid at room temperature.

[2]

- (c) The diagram below shows a distillation column used to separate petroleum into different fractions.



- (i) On the diagram above:

- put a letter **X** to show where the temperature in the column is lowest,
- put a letter **F** to show where the fraction containing the largest molecules is collected,
- put a letter **M** to show where petroleum enters the distillation column.

[3]

- (ii) The refinery gas fraction contains ethane.
Hydrogen is one of the products formed when ethane is cracked.
Complete the symbol equation for the cracking of ethane.



[2]

- (iii) State the conditions needed for cracking.

..... [2]

[Total: 14]

7 Gallium and aluminium are in Group III of the Periodic Table.

(a) The melting point of gallium is 30 °C.

Use the kinetic particle theory to explain what happens when a spoon made of gallium is put into a cup of tea at 40 °C.

In your answer, refer to:

- the change of state which occurs,
- the change in the arrangement of the particles,
- the change in the motion of the particles.

.....

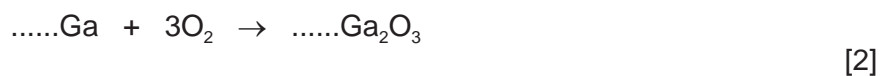
.....

.....

.....

..... [4]

(b) Gallium burns in air at a high temperature to form gallium(III) oxide.
Complete the symbol equation for this reaction.

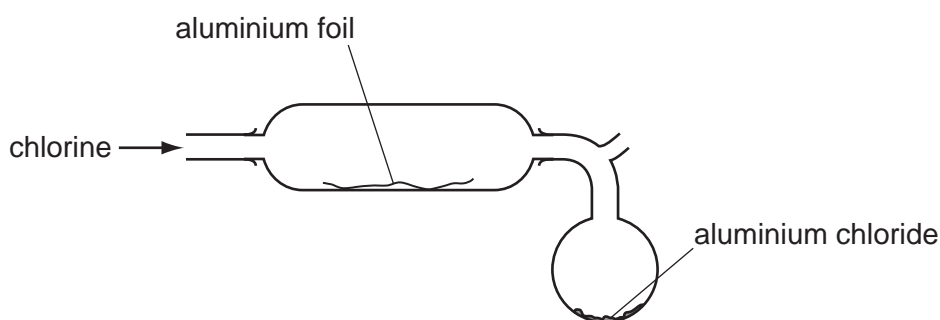


(c) Explain why aluminium is often used in containers for food and drinks.

.....

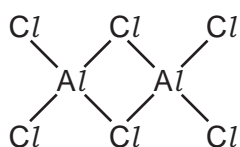
..... [2]

(d) Aluminium chloride can be made by heating aluminium foil in a stream of chlorine.



(i) On the diagram above, draw an arrow to show where heat should be applied. [1]

(ii) At temperatures between 178 °C and 400 °C, aluminium chloride has the structure shown below.



Deduce the molecular formula of this structure.

..... [1]

(iii) Some properties of aluminium and silver are shown in the table below.

	cost	density in g/cm ³	electrical conductivity	melting point /°C
aluminium	high	2.7	good	660
silver	very high	10.5	very good	962

Use the information in the table to suggest why aluminium rather than silver is used in overhead power cables.

..... [1]

[Total: 11]

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DATA SHEET
The Periodic Table of the Elements

		Group																																																	
I	II	III	IV	V	VI	VII	0																																												
1 H Hydrogen 1											2 He Helium 2																																								
3 Li Lithium 3	4 Be Beryllium 4	5 B Boron 5	6 C Carbon 6	7 N Nitrogen 7	8 O Oxygen 8	9 F Fluorine 9	10 Ne Neon 10	11 B Boron 11	12 C Carbon 12	13 Al Aluminium 13	14 Si Silicon 14	15 P Phosphorus 15	16 S Sulfur 16	17 Cl Chlorine 17	18 Ar Argon 18																																				
19 K Potassium 19	20 Ca Calcium 20	21 Sc Scandium 21	22 Ti Titanium 22	23 V Vanadium 23	24 Cr Chromium 24	25 Mn Manganese 25	26 Fe Iron 26	27 Co Cobalt 27	28 Ni Nickel 28	29 Cu Copper 29	30 Zn Zinc 30	31 Ga Gallium 31	32 Ge Germanium 32	33 As Arsenic 33	34 Se Selenium 34	35 Br Bromine 35	36 Kr Krypton 36																																		
37 Rb Rubidium 37	38 Sr Strontium 38	39 Y Yttrium 39	40 Zr Zirconium 40	41 Nb Niobium 41	42 Mo Molybdenum 42	43 Tc Technetium 43	44 Ru Ruthenium 44	45 Rh Rhodium 45	46 Pd Palladium 46	47 Ag Silver 47	48 Cd Cadmium 48	49 In Indium 49	50 Sn Tin 50	51 Sb Antimony 51	52 Te Tellurium 52	53 I Iodine 53	54 Xe Xenon 54																																		
55 Cs Caesium 55	56 Ba Barium 56	57 La Lanthanum 57	72 Hf Hafnium 72	73 Ta Tantalum 73	74 W Tungsten 74	75 Re Rhenium 75	76 Os Osmium 76	77 Ir Iridium 77	78 Pt Platinum 78	79 Au Gold 79	80 Hg Mercury 80	81 Tl Thallium 81	82 Pb Lead 82	83 Bi Bismuth 83	84 Po Polonium 84	85 At Astatine 85	86 Rn Radon 86																																		
87 Fr Francium 87	88 Ra Radium 88	89 Ac Actinium 89																																																	
		*58-71 Lanthanoid series										†90-103 Actinoid series																																							
		<table border="1"> <tr> <td>a</td> <td>X</td> </tr> <tr> <td>b</td> <td></td> </tr> </table>										a	X	b		<table border="1"> <tr> <td>140 Ce Cerium 58</td> <td>141 Pr Praseodymium 59</td> <td>144 Nd Neodymium 60</td> <td>150 Sm Samarium 62</td> <td>152 Eu Europium 63</td> <td>157 Gd Gadolinium 64</td> <td>162 Dy Dysprosium 66</td> <td>165 Ho Holmium 67</td> <td>167 Er Erbium 68</td> <td>169 Tm Thulium 69</td> <td>173 Yb Ytterbium 70</td> <td>175 Lu Lutetium 71</td> </tr> <tr> <td>232 Th Thorium 90</td> <td>238 U Uranium 92</td> <td>238 Pa Protactinium 91</td> <td>238 Np Neptunium 93</td> <td>238 Pu Plutonium 94</td> <td>238 Am Americium 95</td> <td>238 Cm Curium 96</td> <td>238 Bk Berkelium 97</td> <td>238 Cf Californium 98</td> <td>238 Es Einsteinium 99</td> <td>238 Fm Fermium 100</td> <td>238 Md Mendelevium 101</td> <td>238 No Nobelium 102</td> <td>238 Lr Lawrencium 103</td> </tr> </table>										140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71	232 Th Thorium 90	238 U Uranium 92	238 Pa Protactinium 91	238 Np Neptunium 93	238 Pu Plutonium 94	238 Am Americium 95	238 Cm Curium 96	238 Bk Berkelium 97	238 Cf Californium 98	238 Es Einsteinium 99	238 Fm Fermium 100	238 Md Mendelevium 101	238 No Nobelium 102	238 Lr Lawrencium 103
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The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

Key
a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

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