



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CANDIDATE
NAME

CENTRE
NUMBER

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CANDIDATE
NUMBER

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CHEMISTRY

0620/23

Paper 2

October/November 2013

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name in the spaces at the top of this page.

Write in dark blue or black pen.

You may need to use a pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 16.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

This document consists of **15** printed pages and **1** blank page.



1 (a) Choose from the list of metals below to answer the following questions.

aluminium
barium
calcium
iron
lithium
silver

For
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Use

Each metal can be used once, more than once or not at all.

(i) Which metal has an atom with three electrons in its outer electron shell?

..... [1]

(ii) Which **two** metals are in the same Period of the Periodic Table?

..... and [1]

(iii) Which metal has an atom with three protons in its nucleus?

..... [1]

(iv) Which metal has a nitrate which is used to test for halide ions?

..... [1]

(v) Which metal is used in food containers because of its resistance to corrosion?

..... [1]

(b) Describe **two** chemical properties of iron.

1.

2. [2]

(c) Describe briefly how iron from the blast furnace is made into steel.

.....

.....

..... [2]

[Total: 9]

2 Helium is in Group 0 of the Periodic Table.

(a) Describe the structure of a helium atom. Use your Periodic Table to help you.
In your answer, include

- the type and number of subatomic particles present,
- the position of these particles in the atom,
- the relative charges on the particles.

.....

.....

.....

.....

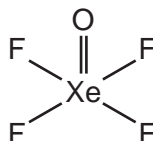
.....

..... [5]

(b) Give **one** use of helium.

..... [1]

(c) Some elements in Group 0 can form compounds with fluorine and oxygen.
The structure of one of these compounds is shown below.



Calculate the relative molecular mass of this compound.
Use your Periodic Table to help you.
You must show all your working.

[2]

(d) Fluorine is a diatomic molecule. It melts at -220°C and boils at -188°C .

(i) What is the physical state of fluorine

at room temperature,

at -200°C ? [2]

(ii) What is meant by the term *diatomic*?

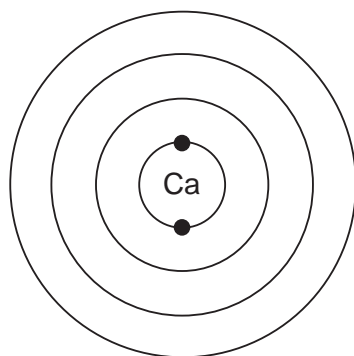
..... [1]

[Total: 11]

3 This question is about calcium and some calcium compounds.

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- (a) Calcium is in Group II of the Periodic Table.
Complete the diagram below to show the electronic structure of calcium.

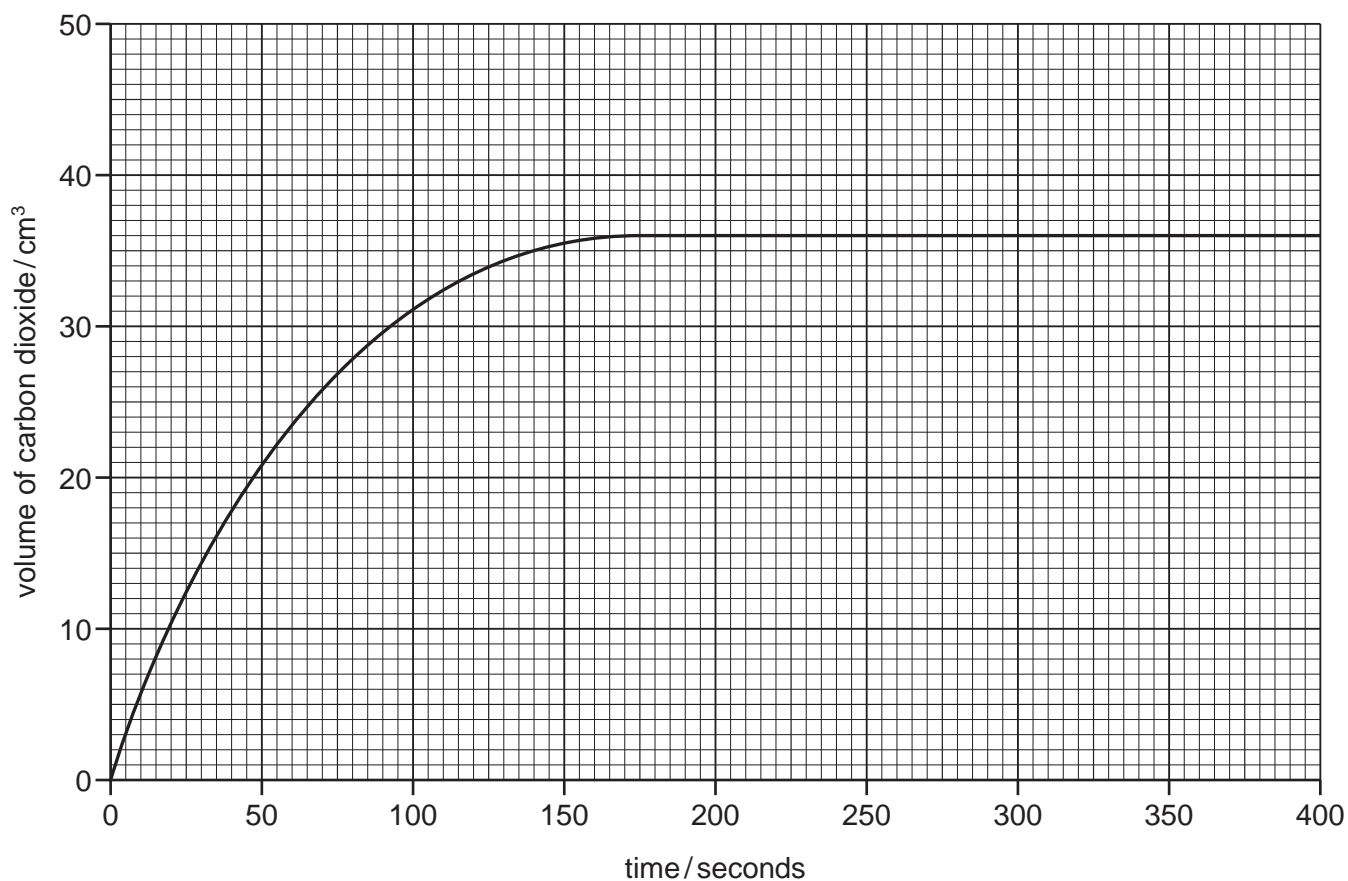


[2]

- (b) Calcium reacts with hydrochloric acid to form a salt with the formula CaCl_2 .
State the name of this salt.

..... [1]

- (c) Calcium carbonate reacts with hydrochloric acid.
The course of this reaction can be followed by measuring the volume of carbon dioxide given off at various time intervals.
The graph below shows the results obtained from an experiment using 0.15 g of calcium carbonate in small pieces.

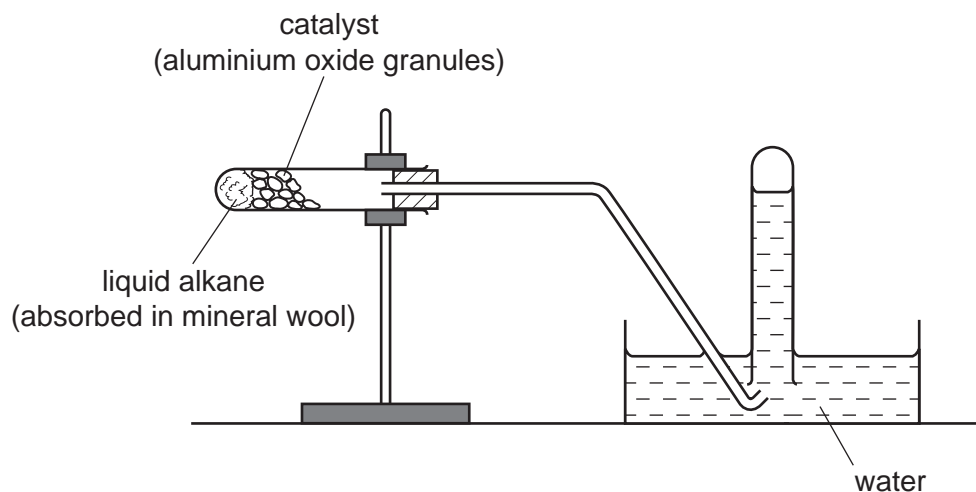


- (i) What volume of gas is given off in the first 75 seconds of the reaction?
..... [1]
- (ii) On the grid opposite, sketch the line you would expect for the same reaction using large pieces of calcium carbonate. Assume that the mass of the calcium carbonate and all other conditions remain the same. [2]
- (iii) What would happen to the rate of this reaction if:
the temperature is increased,
.....
the concentration of hydrochloric acid is decreased?
..... [2]
- (d) When calcium carbonate is heated at high temperatures, calcium oxide and carbon dioxide are formed.
- (i) Which **one** of the following words best describes this reaction?
Put a ring around the correct answer.
combustion decomposition exothermic reduction [1]
- (ii) Describe a test for carbon dioxide.
test
result [2]
- (e) Calcium oxide can be used to neutralise acidic industrial waste.
- (i) Complete the word equation for the reaction of calcium oxide with nitric acid.
calcium oxide + nitric acid → +
..... [2]
- (ii) State **one** other use of calcium oxide.
..... [1]
- (iii) When calcium oxide reacts with water, heat is given off.
State the name given to a chemical reaction which gives off heat.
..... [1]

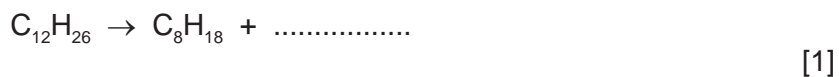
[Total: 15]

- 4 The diagram shows how a liquid alkane can be cracked in a school laboratory to form a mixture of gaseous and liquid hydrocarbons.

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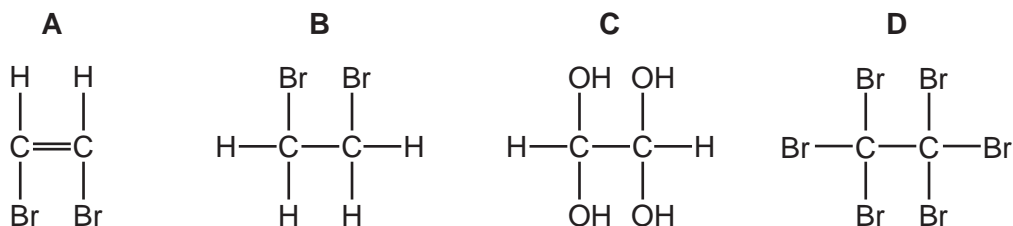
- (a) What piece of apparatus is missing from the diagram?
..... [1]
- (b) On the diagram above, put an **X** to show where the gas is collected. [1]
- (c) What is the purpose of the catalyst?
..... [1]
- (d) Complete the equation to show the cracking of dodecane, $C_{12}H_{26}$, to form octane and **one** other substance.



- (e) Cracking produces a mixture of shorter-chain alkanes and alkenes.
- (i) Describe what you would observe when a few drops of bromine water are added to an alkene.

..... [1]

- (ii) Which one of the following compounds, **A**, **B**, **C** or **D**, is formed when bromine water reacts with ethene?



..... [1]

- (iii) Poly(ethene) is made by combining ethene monomers.
Which one of the following describes this reaction?
Tick **one** box.

decomposition

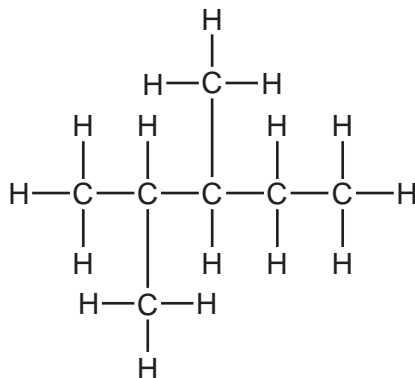
neutralisation

oxidation

polymerisation

[1]

- (f) Many alkanes found in petrol are branched hydrocarbons.
One example is shown below.



- (i) Write the molecular formula for this hydrocarbon.

..... [1]

- (ii) What is meant by the term *hydrocarbon*?

..... [1]

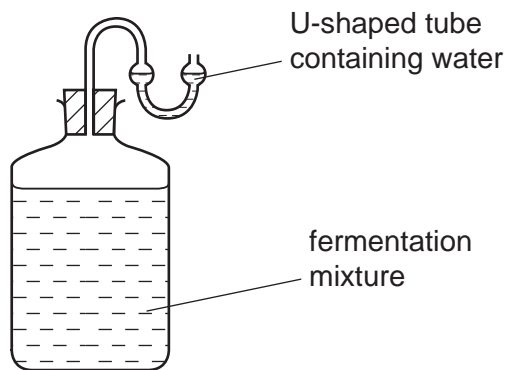
- (g) State the name of the **two** products formed when a hydrocarbon burns in excess air.

..... and [2]

[Total: 11]

5 Ethanol can be made by fermentation.

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(a) Apart from yeast, what other substances are present in the reaction mixture?
Tick **two** boxes.

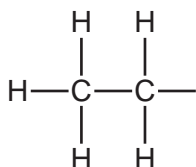
- copper sulfate
- ethene
- sugar
- methane
- water

[2]

(b) What method is used to separate ethanol from the rest of the reaction mixture?

..... [1]

(c) Complete the structure of ethanol.



[1]

(d) Ethanol belongs to the alcohol homologous series.

Which **one** of the following compounds also belongs to the alcohol homologous series?
Put a ring around the correct answer.

butene **hexane** **ethanoic acid** **octanol**

[1]

(e) Describe **one** other way, apart from fermentation, by which ethanol can be made on an industrial scale. Include the necessary reaction conditions in your answer.

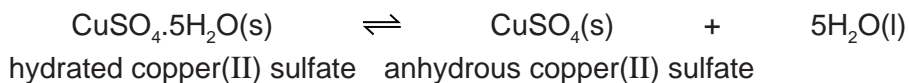
.....

.....

..... [3]

[Total: 8]

- 6 (a) When hydrated copper(II) sulfate is heated, the following reaction occurs:



- (i) What does the sign \rightleftharpoons mean?

..... [1]

- (ii) Explain how this reaction is used as a chemical test for water.

.....

..... [2]

- (iii) Copper(II) sulfate is a salt.

Sodium chloride is also a salt. Solid sodium chloride does not conduct electricity. Suggest **two** things you could do to make solid sodium chloride conduct electricity.

1.

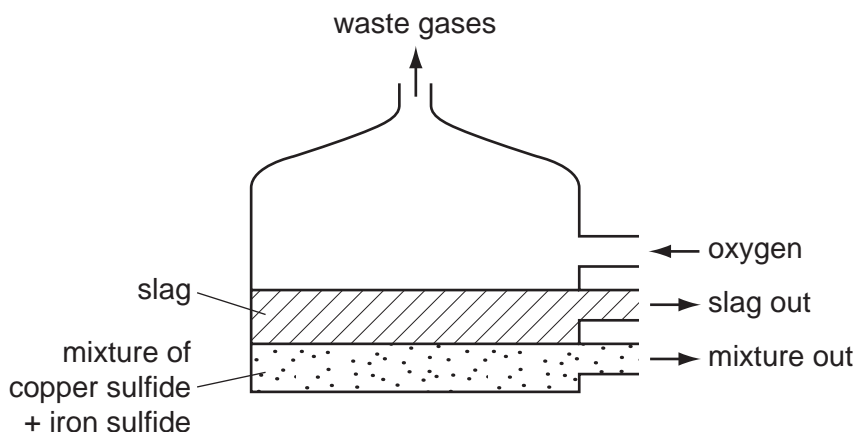
2. [2]

- (b) Copper ore contains copper, iron and sulfur.

Copper is extracted by heating copper ore with sand and oxygen.

- (i) In the first stage of this process, the copper ore is heated in a furnace.

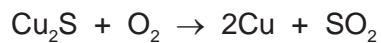
A liquid mixture containing copper sulfide and iron sulfide is formed. The sand reacts with the impurities to form a slag.



What information in the diagram above suggests that the slag is less dense than the mixture of copper and iron sulfides.

..... [1]

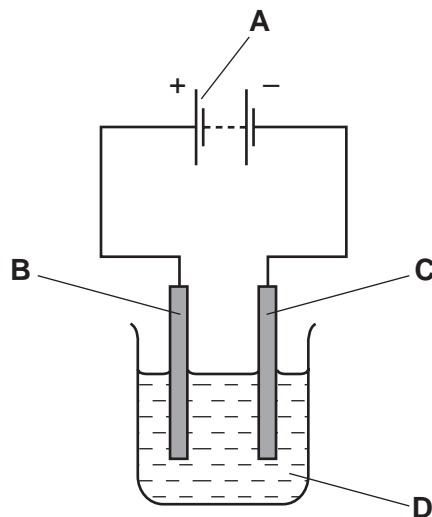
- (ii) In a later stage, copper sulfide is reacted with more oxygen.



How does this equation show that the sulfur in copper sulfide gets oxidised?

..... [1]

- (iii) Copper is purified by electrolysis using copper electrodes.



Which letter, **A**, **B**, **C** or **D**, in the diagram above represents

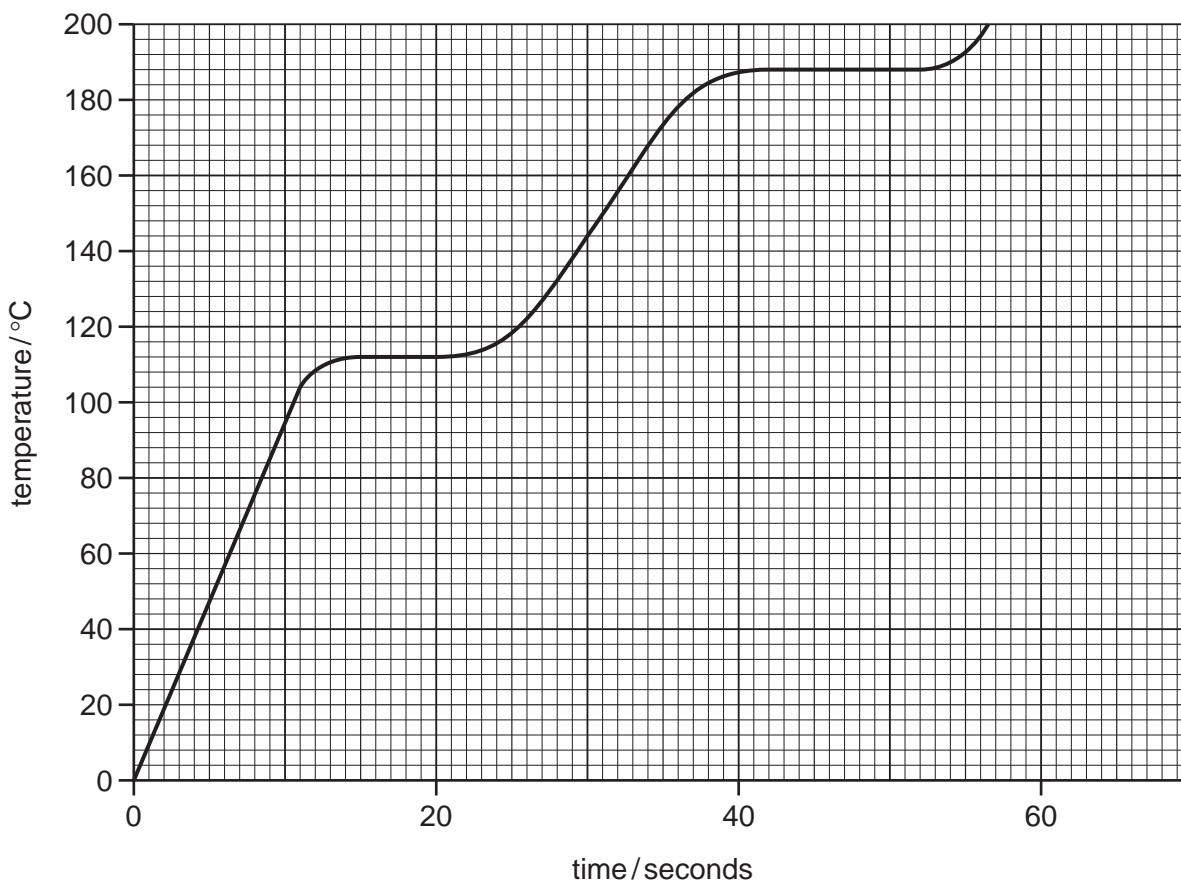
the cathode,

the electrolyte? [2]

[Total: 9]

- 7 The graph below shows how the temperature rises with time when a solid, **P**, is heated steadily and changes to a liquid and then to a gas.

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- (a) Use the information on the graph to deduce

the melting point of **P**,

the state of **P** at 160 °C. [2]

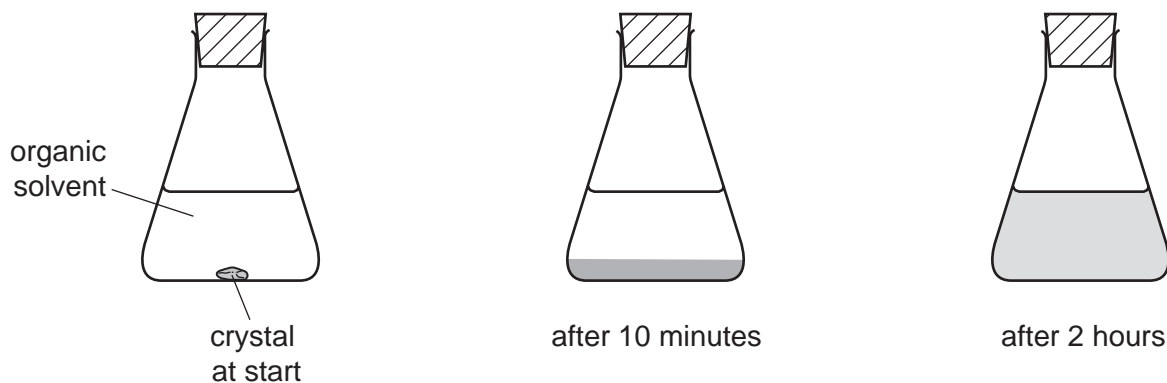
- (b) Explain what happens to the arrangement and motion of the particles when a solid changes to a liquid.

arrangement

motion [2]

- (c) A student placed a purple crystal in a flask of organic solvent.
After 10 minutes, the crystal had completely disappeared and a dense purple colour was observed at the bottom of the flask.
After 2 hours, the purple colour had spread throughout the solvent.

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Use the kinetic particle theory to explain these observations.

.....

.....

.....

..... [3]

[Total: 7]

- 8 (a) State **two** differences between a mixture and a compound.

.....

.....

..... [2]

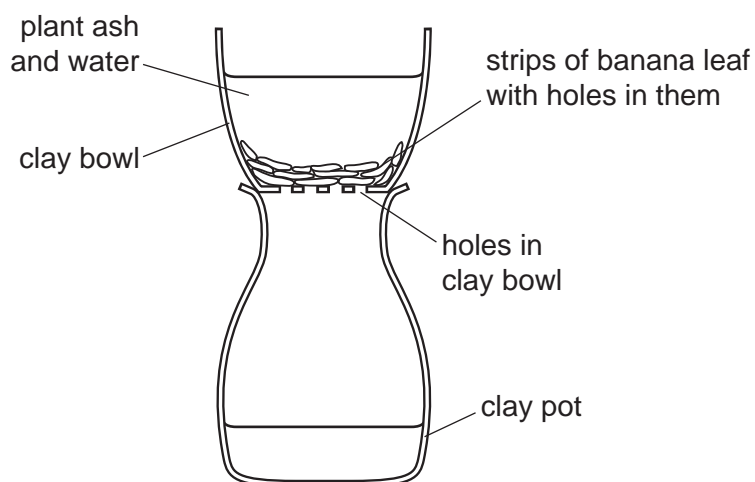
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- (b) Plant ash is a mixture of large insoluble particles and salts which are soluble in water.

In parts of Africa, salts are traditionally obtained from plant ash.

Water is added to the plant ash.

The apparatus shown below is then used to remove the insoluble particles.



Explain how this apparatus separates the salts from the insoluble particles.

.....

.....

.....

..... [2]

- (c) The composition and solubility of some salts found in the ash from the papyrus plant are shown in the table below.

salt	ion present in the salt	mass of salt per 100g of ash/g	solubility of salt in g/dm ³
magnesium sulfate	Mg ²⁺ and SO ₄ ²⁻	5	220
potassium carbonate	K ⁺ and CO ₃ ²⁻	10	1120
potassium chloride	K ⁺ and Cl ⁻	18	359
potassium sulfate		4	122
sodium carbonate	Na ⁺ and CO ₃ ²⁻	12	70
sodium chloride	Na ⁺ and Cl ⁻	40	359

- (i) Which salt in the table has the lowest solubility in g/dm³?

..... [1]

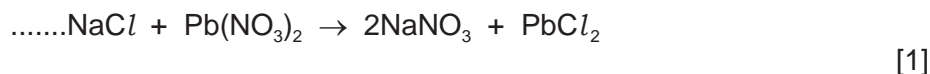
- (ii) Which negatively-charged ion is present in the highest amount in the ash?

..... [1]

- (iii) Write the symbols for the **two** ions present in potassium sulfate.

..... [2]

- (d) Sodium chloride reacts with lead(II) nitrate to form sodium nitrate and lead(II) chloride. Complete the symbol equation for this reaction.



- (e) Complete the following sentence about the formation of chloride ions.

Chloride ions are formed when chlorine atoms gain [1]

[Total: 10]

DATA SHEET
The Periodic Table of the Elements

		Group																		
I	II	III	IV	V	VI	VII	0													
1 H Hydrogen 1											2 He Helium 2									
3 Li Lithium 3	4 Be Beryllium 4	5 B Boron 5	6 C Carbon 6	7 N Nitrogen 7	8 O Oxygen 8	9 F Fluorine 9	10 Ne Neon 10	11 B Boron 11	12 C Carbon 12	13 Al Aluminium 13	14 Si Silicon 14	15 P Phosphorus 15	16 S Sulfur 16	17 Cl Chlorine 17	18 Ar Argon 18					
19 K Potassium 19	20 Ca Calcium 20	21 Sc Scandium 21	22 Ti Titanium 22	23 V Vanadium 23	24 Cr Chromium 24	25 Mn Manganese 25	26 Fe Iron 26	27 Co Cobalt 27	28 Ni Nickel 28	29 Cu Copper 29	30 Zn Zinc 30	31 Ga Gallium 31	32 Ge Germanium 32	33 As Arsenic 33	34 Se Selenium 34	35 Br Bromine 35	36 Kr Krypton 36			
37 Rb Rubidium 37	38 Sr Strontium 38	39 Y Yttrium 39	40 Zr Zirconium 40	41 Nb Niobium 41	42 Mo Molybdenum 42	43 Tc Technetium 43	44 Ru Ruthenium 44	45 Rh Rhodium 45	46 Pd Palladium 46	47 Ag Silver 47	48 Cd Cadmium 48	49 In Indium 49	50 Sn Tin 50	51 Sb Antimony 51	52 Te Tellurium 52	53 I Iodine 53	54 Xe Xenon 54			
55 Cs Caesium 55	56 Ba Barium 56	57 La Lanthanum 57	72 Hf Hafnium 72	73 Ta Tantalum 73	74 W Tungsten 74	75 Re Rhenium 75	76 Os Osmium 76	77 Ir Iridium 77	78 Pt Platinum 78	79 Au Gold 79	80 Hg Mercury 80	81 Tl Thallium 81	82 Pb Lead 82	83 Bi Bismuth 83	84 Po Polonium 84	85 At Astatine 85	86 Rn Radon 86			
87 Fr Francium 87	88 Ra Radium 88	89 Ac Actinium 89																		
*58-71 Lanthanoid series																				
†90-103 Actinoid series																				
<table border="1"> <tr> <td>a</td> <td>X</td> </tr> <tr> <td>b</td> <td></td> </tr> </table> <p>Key a = relative atomic mass X = atomic symbol b = proton (atomic) number</p>																	a	X	b	
a	X																			
b																				
133 Cs Caesium 133	137 Ba Barium 137	138 La Lanthanum 138	139 Ce Cerium 139	140 Pr Praseodymium 140	141 Nd Neodymium 141	142 Pm Promethium 142	143 Sm Samarium 143	144 Eu Europium 144	145 Gd Gadolinium 145	146 Tb Terbium 146	147 Dy Dysprosium 147	148 Ho Holmium 148	149 Er Erbium 149	150 Tm Thulium 150	151 Yb Ytterbium 151	152 Lu Lutetium 152				
90 Th Thorium 90	91 Pa Protactinium 91	92 U Uranium 92	93 Np Neptunium 93	94 Pu Plutonium 94	95 Am Americium 95	96 Cm Curium 96	97 Bk Berkelium 97	98 Cf Californium 98	99 Es Einsteinium 99	100 Fm Fermium 100	101 Md Mendelevium 101	102 No Nobelium 102	103 Lr Lawrencium 103							

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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