#### UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

International General Certificate of Secondary Education

#### MARK SCHEME for the November 2004 question paper

#### 0620 CHEMISTRY

0620/05

Paper 5 (Practical Test), maximum mark 40

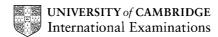
This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which Examiners were initially instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began. Any substantial changes to the mark scheme that arose from these discussions will be recorded in the published *Report on the Examination*.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the *Report on the Examination*.

• CIE will not enter into discussion or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the November 2004 question papers for most IGCSE and GCE Advanced Level syllabuses.



Grade thresholds taken for Syllabus 0620 (Chemistry) in the November 2004 examination.

	maximum mark available	minimum mark required for grade:				
		А	С	Е	F	
Component 5	40	32	27	23	20	

The threshold (minimum mark) for B is set halfway between those for Grades A and C. The threshold (minimum mark) for D is set halfway between those for Grades C and E. The threshold (minimum mark) for G is set as many marks below the F threshold as the E threshold is above it.

Grade A\* does not exist at the level of an individual component.



PMT

November 2004

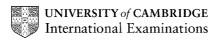
## **INTERNATIONAL GCSE**

# MARK SCHEME

# MAXIMUM MARK: 40

## SYLLABUS/COMPONENT: 0620/05

CHEMISTRY Practical Test



	Page 1	Mark Scheme			Paper		
		IGCSE – November 2	004	0620	5		
1	Table of results	ble of results					
	initial temperatures co	ompleted (1)	final temperatures completed (1)				
	decreasing (1)	decreasing (1)		comparable to supervisors (2) (5)			
	(a) Graph. Points co	(a) Graph. Points correctly plotted (3),		-1 for each incorrect			
	straight line (1)	t line (1)			(4)		
	(b) (i) temperature for 3.5 g from graph (1)		indication (1)		(2)		
	°C (1)				(1)		
	(ii) temperature fo	or 6 g from graph(1)	extrapolation (1) (2)				
	(c) endothermic (1)				(1)		
	(d) lower temperature	(d) lower temperature changes (1)		more dilute/water to heat up (1)			
	(e) larger surface are	(e) larger surface area (1)		dissolves faster/easier (1)			
	(f) initial temperature	from table $\pm$ 2 °C (1)	room temperature/reaction finished (1)(2)				
	(g) change to apparatus e.g. insulation/burette/pipette etc. (1)						
				•	sub total 22		
2	(a) white (1)	powder/crystals (1)	not precipitate		(2)		
	(b) indicator paper go	<b>b)</b> indicator paper goes blue (1) red (1) smell/solid is smaller (1) <b>n</b>					
	white deposit (1)	deposit (1) on tube e.g. top/side (1)					
	<b>(c)</b> pH 4-6 (1)				(1)		
	(d) (i) white (1)		precipitate (1)		(2)		
	(ii) white (1)		precipitate (1)		(2)		
	(iii) pH greater tha	an 7 (1)	reference to smell/colour of indicator paper (1)		dicator (2)		
	(e) ammonia (1)				(1)		
	<b>(f)</b> alkaline gas (1)	ammonia named (2)	sublimation (1)		max 2 (2)		
	<b>(g)</b> ammonium (1)	chloride (1)			(2)		
				:	sub total 18		
					Total 40		

Total 40