

Please check the examination details below before entering your candidate information

Candidate surname					Other names				
Centre Number				Candidate Number					
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Pearson Edexcel International GCSE (9–1)

Time 1 hour 15 minutes

Paper reference **4CH1/2C**

Chemistry

Unit: 4CH1

PAPER: 2C

You must have:
Calculator, ruler

Total Marks

Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*
- Show all the steps in any calculations and state the units.

Information

- The total mark for this paper is 70.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*

Advice

- Read each question carefully before you start to answer it.
- Write your answers neatly and in good English.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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The Periodic Table of the Elements

1	2	3	4	5	6	7	0																										
7 Li lithium 3	9 Be beryllium 4	11 Na sodium 11	12 C carbon 6	13 Al aluminium 13	14 N nitrogen 7	15 P phosphorus 15	16 O oxygen 8	17 F fluorine 9	18 Ne neon 10																								
19 K potassium 19	20 Ca calcium 20	23 Sc scandium 21	24 Ti titanium 22	25 V vanadium 23	26 Cr chromium 24	27 Mn manganese 25	28 Fe iron 26	29 Co cobalt 27	30 Ni nickel 28	31 Cu copper 29	32 Zn zinc 30	33 Ga gallium 31	34 Ge germanium 32	35 As arsenic 33	36 Se selenium 34	37 Rb rubidium 37	38 Sr strontium 38	39 Y yttrium 39	40 Zr zirconium 40	41 Nb niobium 41	42 Mo molybdenum 42	43 Tc technetium 43	44 Ru ruthenium 44	45 Rh rhodium 45	46 Pd palladium 46	47 Ag silver 47	48 Cd cadmium 48	49 In indium 49	50 Sn tin 50	51 Sb antimony 51	52 Te tellurium 52	53 I iodine 53	54 Xe xenon 54
55 Cs caesium 55	56 Ba barium 56	57 La* lanthanum 57	58 Ce cerium 58	59 Pr praseodymium 59	60 Nd neodymium 60	61 Pm promethium 61	62 Sm samarium 62	63 Eu europium 63	64 Gd gadolinium 64	65 Tb terbium 65	66 Dy dysprosium 66	67 Ho holmium 67	68 Er erbium 68	69 Tm thulium 69	70 Yb ytterbium 70	71 Lu lutetium 71	72 Hf hafnium 72	73 Ta tantalum 73	74 W tungsten 74	75 Re rhenium 75	76 Os osmium 76	77 Ir iridium 77	78 Pt platinum 78	79 Au gold 79	80 Hg mercury 80	81 Tl thallium 81	82 Pb lead 82	83 Bi bismuth 83	84 Po polonium 84	85 At astatine 85	86 Rn radon 86		
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[277] Hs hassium 108	[268] Mt meitnerium 109	[271] Ds darmstadtium 110	[272] Rg roentgenium 111	Elements with atomic numbers 112–116 have been reported but not fully authenticated						[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86														

1	H hydrogen 1
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relative atomic mass
atomic symbol
name
atomic (proton) number

* The lanthanoids (atomic numbers 58–71) and the actinoids (atomic numbers 90–103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.



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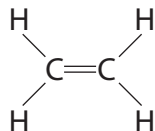


Answer ALL questions.

Some questions must be answered with a cross in a box . If you change your mind about an answer, put a line through the box and then mark your new answer with a cross .

1 This question is about the unsaturated hydrocarbon, ethene.

The displayed formula of ethene is



(a) (i) State the meaning of the term **hydrocarbon**.

(2)

.....

.....

.....

.....

(ii) Give the reason why ethene is described as unsaturated.

(1)

.....

.....

(b) Ethene is bubbled through bromine water until there is no further colour change.

Which of these is the appearance of the solution formed?

(1)

- A** colourless
- B** orange
- C** purple
- D** red

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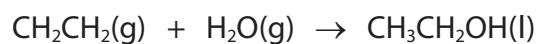
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(c) Ethanol is produced industrially by the reaction between ethene and steam.

The equation for the reaction is



(i) State the temperature and pressure used in this reaction.

(2)

temperature

pressure

(ii) Give the **molecular** formula of ethanol.

(1)

(Total for Question 1 = 7 marks)

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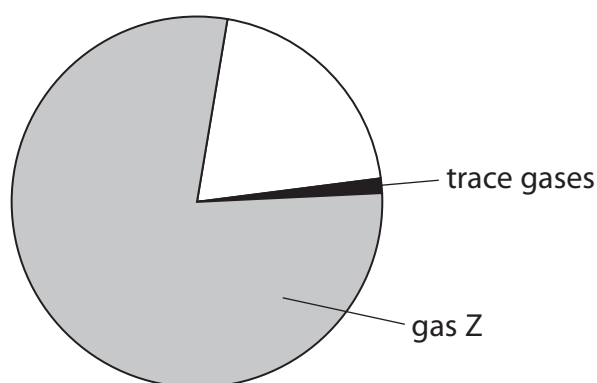
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2 This question is about gases in the air.

The pie chart represents the percentages of gases in dry, unpolluted air.

Gases with percentages of less than 1% in air are called trace gases.



(a) (i) Which of these is gas Z?

(1)

- A hydrogen
- B methane
- C neon
- D nitrogen

(ii) Which of these is the approximate percentage of oxygen in dry, unpolluted air?

(1)

- A 0.04%
- B 0.9%
- C 21%
- D 35%

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(b) One of the trace gases is carbon dioxide.

- (i) Identify **two** reactions that produce carbon dioxide by placing a tick (✓) in two boxes.

(2)

cracking an alkane	
complete combustion of an alkane	
reaction between magnesium and hydrochloric acid	
rusting of iron	
thermal decomposition of copper(II) carbonate	

- (ii) Name an environmental problem that is caused by the percentage of carbon dioxide increasing in the atmosphere.

(1)

- (iii) Name the trace gas with the highest percentage in dry, unpolluted air.

(1)

(c) Rainwater is acidic because carbon dioxide dissolves in water to form carbonic acid.

Acid rain is more acidic than rainwater because acidic pollutant gases also dissolve in water.

- (i) Give the name of the acid that forms when nitrogen dioxide dissolves in water.

(1)

- (ii) Name another pollutant gas that also forms acid rain.

(1)

(Total for Question 2 = 8 marks)

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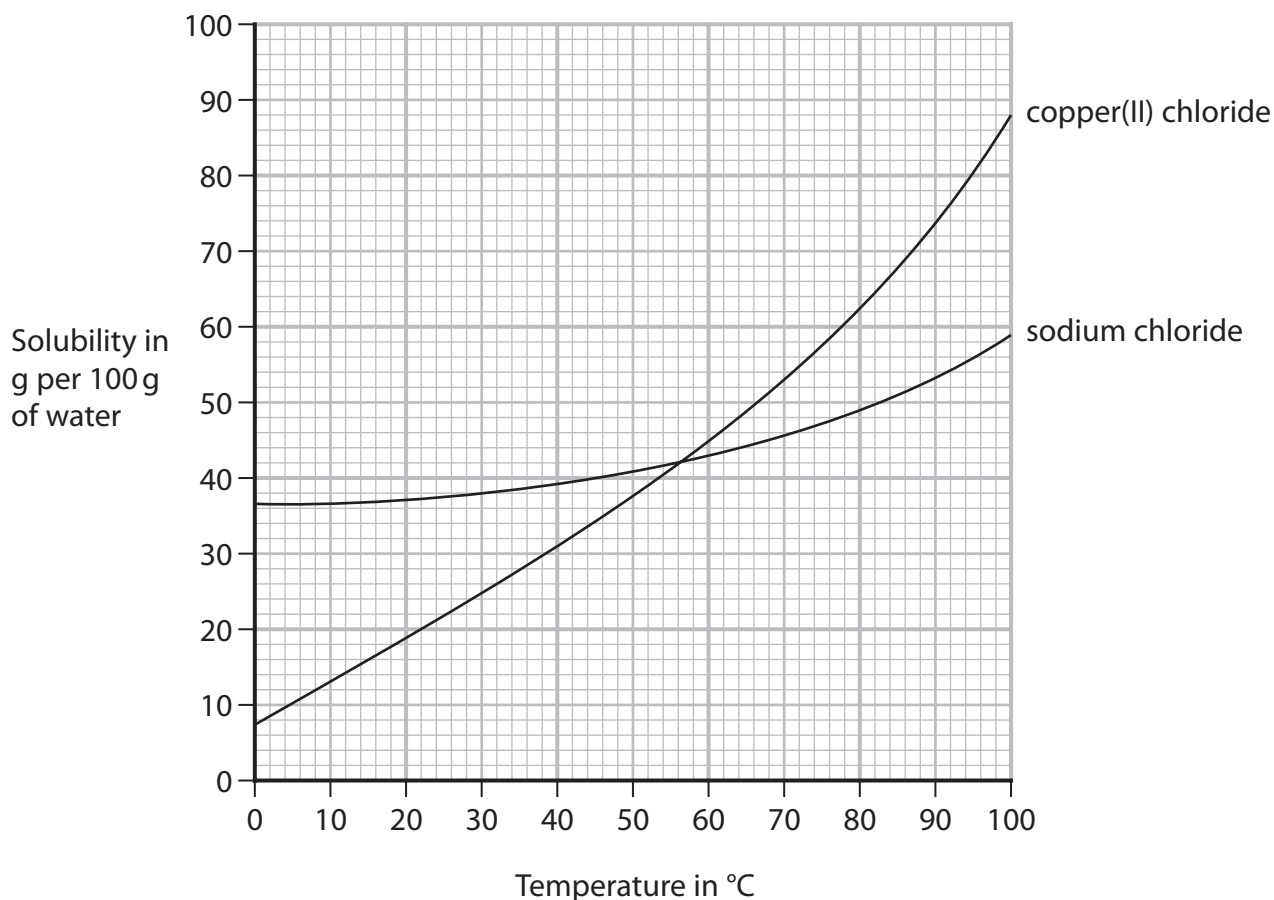
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3 This question is about solubility.

(a) The graph shows the solubilities of copper(II) chloride and sodium chloride at different temperatures.



(i) Determine the temperature at which copper(II) chloride and sodium chloride have the same solubility.

Show on the graph how you obtained your answer.

(2)

temperature = °C

(ii) A saturated solution of copper(II) chloride in 100 g of water is cooled from 40 °C to 10 °C.

Determine the mass, in grams, of copper(II) chloride that crystallises.

(2)

mass of copper(II) chloride = g



P 7 0 7 0 2 A 0 9 2 8

(b) A student uses this method to determine the solubility of potassium chloride in water at room temperature.

- record the mass of an empty evaporating basin
- pour some saturated potassium chloride solution into the evaporating basin
- record the mass of the evaporating basin and saturated potassium chloride solution
- heat the evaporating basin to remove all the water
- record the mass of the evaporating basin and the dry potassium chloride

The table shows the student's results.

	Mass in grams
evaporating basin	58.1
evaporating basin and saturated potassium chloride solution	78.2
evaporating basin and dry potassium chloride	63.2

(i) Calculate the mass of dry potassium chloride obtained.

(1)

mass = g

(ii) Calculate the mass of water removed.

(1)

mass = g



(iii) Calculate the solubility of potassium chloride in grams per 100 grams of water.

(2)

solubility = g per 100 g of water

(iv) Suggest why the student's method is **not** suitable for determining the solubility of hydrated copper(II) sulfate.

(1)

(Total for Question 3 = 9 marks)

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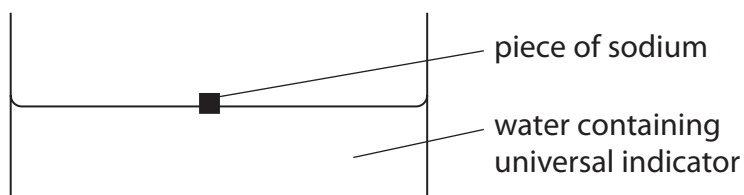
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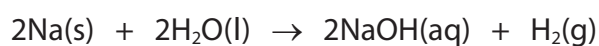
P 7 0 7 0 2 A 0 1 1 2 8

4 This question is about the reactions of Group 1 metals with water.

(a) A teacher adds a piece of sodium to some water containing universal indicator.



The equation for this reaction is



The sodium floats on the surface of the water and the universal indicator changes colour because an alkaline solution is formed.

(i) Give two other observations.

(2)

1

.....

2

.....

(ii) Give the final colour of the universal indicator.

(1)

.....

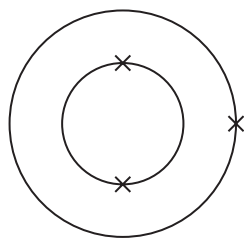
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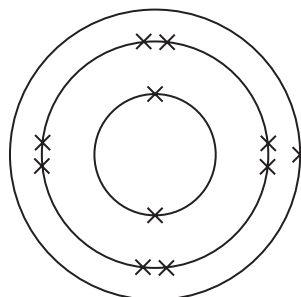
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(b) The diagram represents an atom of lithium and an atom of sodium.



lithium



sodium

(i) Give a reason why lithium and sodium have similar reactions with water.

(1)

(ii) Explain why lithium is less reactive than sodium.

(3)

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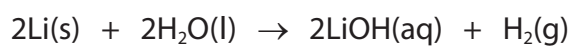
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- (c) The teacher adds 0.150 g of lithium to an excess of water and collects the hydrogen gas produced.

The equation for the reaction is



The teacher collects 254 cm³ of hydrogen gas at room temperature and pressure (rtp).

Show by calculation that 1 mol of hydrogen gas has a volume of approximately 24 000 cm³ at rtp.

(4)

(Total for Question 4 = 11 marks)



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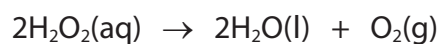
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5 Hydrogen peroxide solution decomposes to give water and oxygen gas.

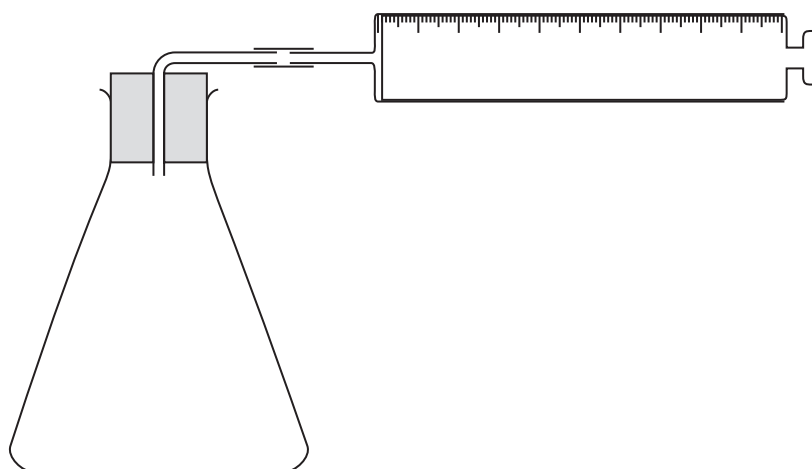
The equation for this reaction is



- (a) Three different solids are catalysts for the decomposition of hydrogen peroxide solution.

A student is given hydrogen peroxide solution and a sample of each of the solid catalysts.

The student has a timer, a measuring cylinder, a balance and the apparatus shown in the diagram.



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Describe a method the student could use to find which of the three solids is the most effective catalyst for the decomposition of hydrogen peroxide solution.

(5)

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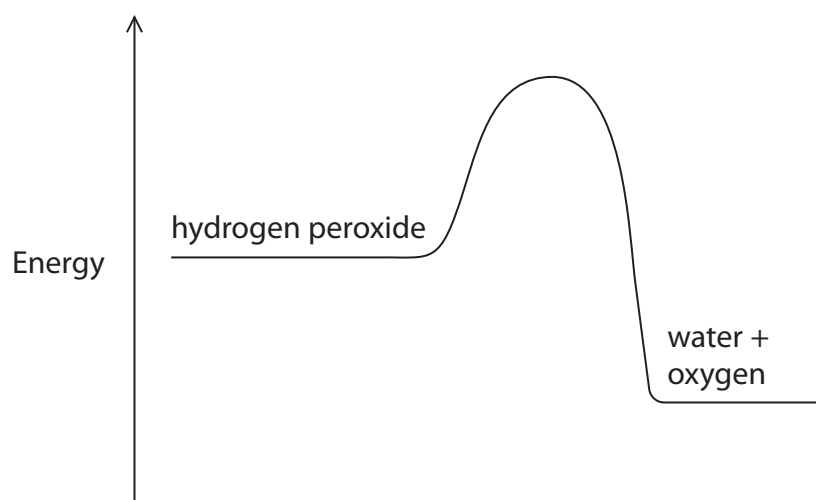
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Area with horizontal dotted lines for writing the answer.



- (b) The diagram shows the reaction profile for the decomposition of hydrogen peroxide without a catalyst.



- (i) Label the diagram to show the activation energy (E_a) and the enthalpy change (ΔH) for this reaction. (2)
- (ii) On the diagram, draw a curve to show the reaction profile for the same reaction when a catalyst is used. (1)

(Total for Question 5 = 8 marks)



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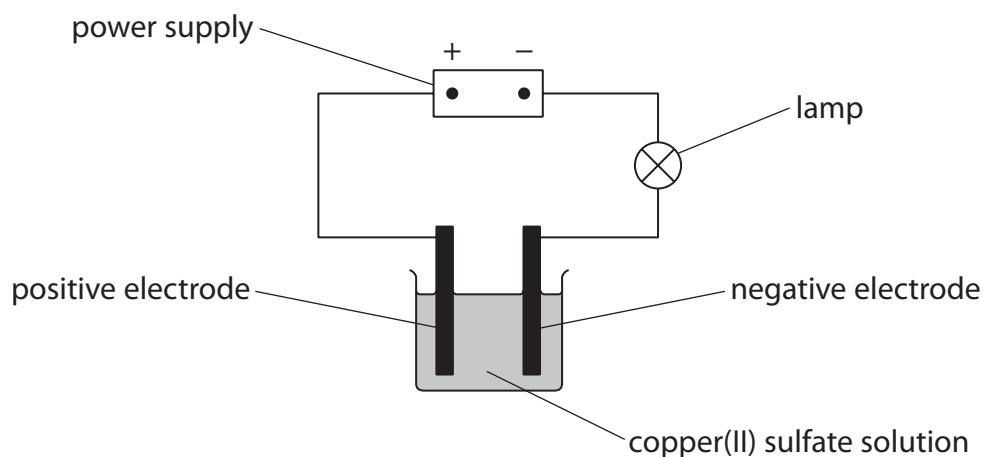
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6 This question is about the electrolysis of copper(II) sulfate solution.

(a) The diagram shows the apparatus used for the electrolysis.



A student records the total increase in mass of the negative electrode every minute for 8 minutes.

The table shows the results.

Time in minutes	Total increase in mass of the negative electrode in grams
0	0.00
1	0.15
2	0.27
3	0.34
4	0.39
5	0.41
6	0.42
7	0.42
8	0.42

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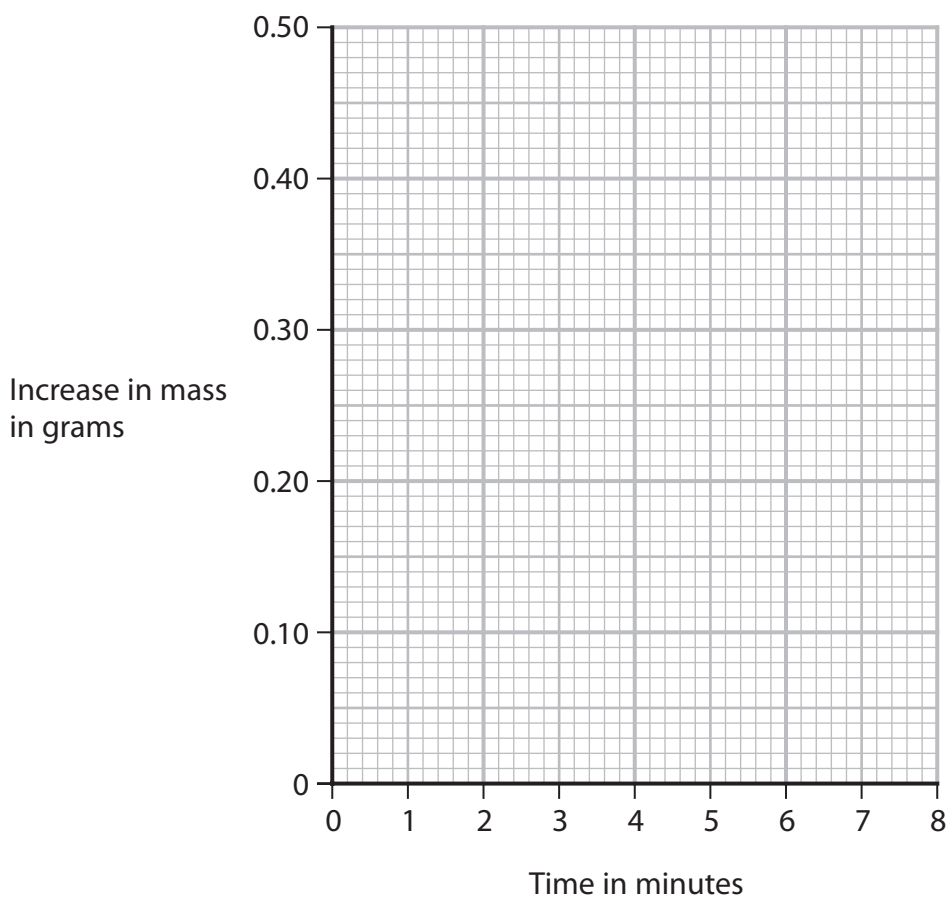


(i) Plot the student's results.

(1)

(ii) Draw a curve of best fit.

(1)



(iii) Explain the shape of the graph.

(2)

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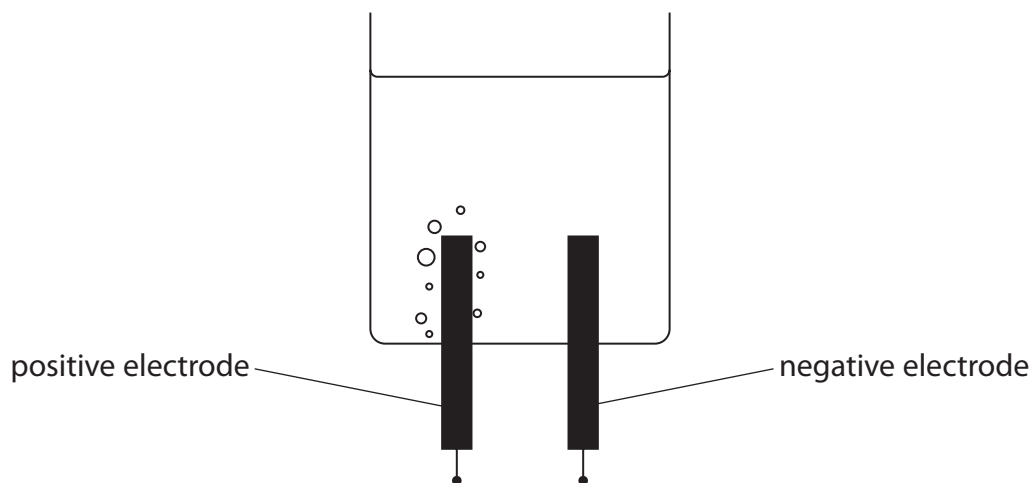
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P 7 0 7 0 2 A 0 2 1 2 8

(b) The product at the positive electrode is oxygen gas.

(i) The student repeats the electrolysis using different apparatus.



Describe how the student should collect a sample of pure oxygen at the positive electrode.

(2)

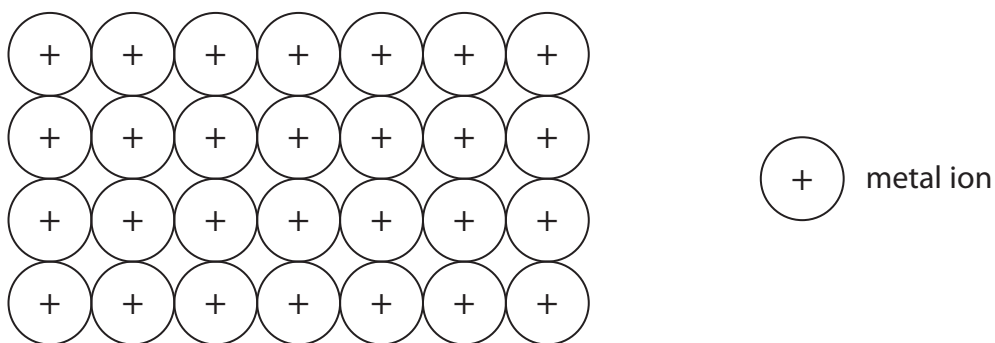
(ii) Give an ionic half-equation for the formation of oxygen.

(2)



- (c) The wire used to connect the power supply to the electrodes is made of copper metal.

The diagram shows the arrangement of the ions in a metal.



- (i) Metals that are malleable can also be stretched to form long, thin wires.

Suggest why metals can be stretched to form wires.

(2)

- (ii) Explain why metals conduct electricity.

(2)

(Total for Question 6 = 12 marks)

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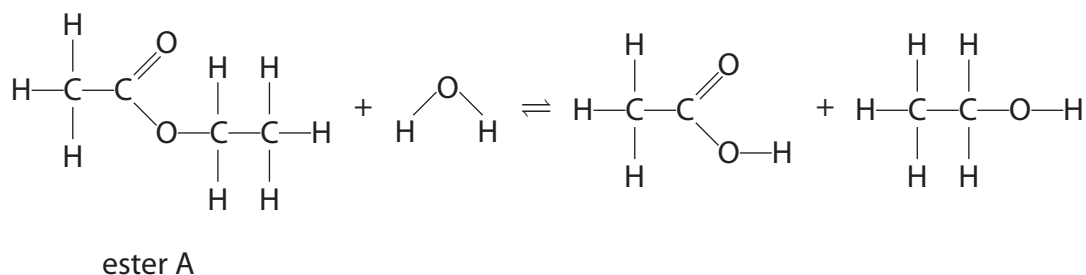


P 7 0 7 0 2 A 0 2 3 2 8

7 This question is about esters.

Ester A reacts with water to form ethanoic acid and ethanol.

The displayed formulae of the reactants and products are shown in this equation



The molar enthalpy change (ΔH) for the reaction is 0 kJ/mol.

(a) (i) Draw a ring around the functional group in ester A. (1)

(ii) Give the name of ester A. (1)

(iii) Describe a chemical test, other than using an indicator, to show that the reaction mixture contains ethanoic acid. (2)

.....

.....

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(b) Explain why the molar enthalpy change (ΔH) for the reaction between ester A and water is 0 kJ/mol.

In your answer, refer to the bonds broken and the bonds formed. (2)

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(c) A mixture of ester A and water is left in a sealed container until the reaction mixture reaches dynamic equilibrium.

(i) Describe what is meant by dynamic equilibrium.

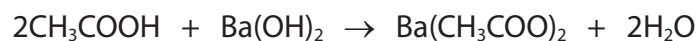
(2)

(ii) Explain why adding a catalyst does not change the position of equilibrium.

(2)

(d) The ethanoic acid produced in the reaction is completely neutralised by 22.75 cm³ of 0.150 mol/dm³ barium hydroxide solution.

The equation for the neutralisation reaction is



Calculate the amount, in moles, of ethanoic acid neutralised.

Give your answer to 3 significant figures.

(3)

amount = mol

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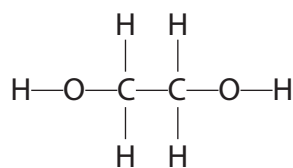
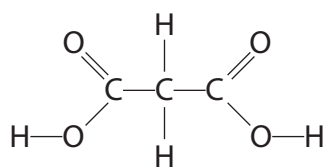
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(e) The structures of two organic compounds are shown.



These compounds react together to form a polymer.

Give the repeat unit of the polymer formed.

(2)

(Total for Question 7 = 15 marks)

TOTAL FOR PAPER = 70 MARKS



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